

## Reviewing Content

42. The solvent is the substance in which the solute is dissolved.
43. Random collisions of the solvent molecules with the solute particles provide enough force to overcome gravity.
44. **solubility**: the amount of a substance that dissolves in a given quantity of solvent at specified conditions of temperature and pressure to produce a saturated solution.  
**saturated solution**: a solution containing the maximum amount of solute for a given amount of solvent at a constant temperature and pressure. **unsaturated solution**: a solution that contains less solute than a saturated solution at a given temperature and pressure. **miscible**: describes liquids that dissolve in each other. **immiscible**: describes liquids that are insoluble in each other.
45. Particles of solute crystallize.
46. No; if there were undissolved solute, the excess solute would come out of a supersaturated solution.
47.  $5.55 \times 10^2$  g  $\text{AgNO}_3$
48. Solubility increases with pressure.
49. a.  $1.6 \times 10^{-2}$  g/L  
b.  $4.7 \times 10^{-2}$  g/L
50. *Dilute* and *concentrated* are relative terms and are not quantitative. Molarity provides the exact number of moles of solute per liter of solution.
51. Molarity is the number of moles of solute dissolved in one liter of solution.  
a. 1.3M KCl  
b.  $3.3 \times 10^{-1}$ M  $\text{MgCl}_2$
52.  $2.00 \times 10^3$  mL
53. a.  $5.0 \times 10^{-3}$  mol NaCl, 29 g NaCl  
b. 1.0 mol  $\text{KNO}_3$ ,  $1.0 \times 10^{-2}$  g  $\text{KNO}_3$   
c.  $2.5 \times 10^3$  mol  $\text{CaCl}_2$ , 2.8 g  $\text{CaCl}_2$
54. a.  $2.3 \times 10^1$  g NaCl  
b. 2.0 g  $\text{MgCl}_2$
55. a. 16% (v/v) ethanol  
b. 63.6% (v/v) isopropyl alcohol
56. Colligative properties are properties of a solution that depend only on the number of solute particles; boiling-point elevation, freezing-point depression, and vapor-pressure lowering. Boiling points are elevated because shells of solvent form around solute particles, reducing the amount of solvent molecules that have sufficient energy to escape the solution; relative to the pure solvent, the amount of energy required to cause vaporization or boiling increases. Solutes disrupt the ordering of the solvent structure, so more kinetic energy must be withdrawn from a solution for it to solidify. This lowers the freezing point of the solution.
57. a. sea water  
b. 1.5M  $\text{KNO}_3$   
c. 0.100M  $\text{MgCl}_2$
58. The effective molality of the  $\text{Ca}(\text{NO}_3)_2$  solution is 3m. The effective molality of the  $\text{NaNO}_3$  solution is 2m.
59. When vapor pressure is lowered relative to pure solvent, more energy must be supplied to reach the boiling point; thus the boiling point is increased relative to pure solvent.
60. The salt lowers the freezing point of the ice-water cooling mixture.
61. 1M solution: 1 mol of solute in 1 L of solution; 1m solution: 1 mol of solute in 1000 g of solvent
62. Add 27.0 g  $\text{H}_2\text{O}$  to 32.0 g  $\text{CH}_3\text{OH}$ .
63. a. 100.26°C  
b. 101.54°C
64. a. -4.46°C  
b. -2.2°C
65. a. -1.1°C  
b. -0.74°C  
c. -1.5°C

## Understanding Concepts

66. a. The freezing-point depression is twice as great for solute B; solute B must provide twice as many particles in solution.

# Prentice Hall Chemistry Answer Key Chapter 16

**Ray Quincy Brewster, William Edwin  
McEwen**



## Prentice Hall Chemistry Answer Key Chapter 16:

*An Introduction to Spectroscopy, Atomic Structure and Chemical Bonding* Terry L. Meek, 1998 An Introduction to Spectroscopy presents the most fundamental concepts of inorganic chemistry at a level appropriate for first year students and in a manner comprehensible to them This is true even of difficult topics such as the wave mechanical atom symmetry elements and symmetry operations and the ligand group orbital approach to bonding The book contains many useful diagrams illustrating among other things the angular dependence of atomic wave functions the derivation of energy level diagrams for polyatomic molecules close packed lattices and ionic crystal structures The diagrams of the periodic variation of atomic and molecular properties showing trends across periods and down groups simultaneously are especially instructive Spectroscopy is presented mainly as a tool for the elucidation of atomic and molecular structures Each chapter begins with a clear and concise statement of What Every First year Student Should Know About outlining the background knowledge that the student is assumed to have from previous courses and thus pointing out what topics might need to be reviewed There are also detailed statements of the objectives of each chapter a number of worked examples interspersed in the text and a comprehensive set of problems and exercises to test the student's understanding Tables of data throughout the text and appendices at the end provide much valuable information

Books in Print, 1954      **Chemistry for Health-related Sciences** Curtis T. Sears, Conrad L. Stanitski, 1983      **Solutions Manual with Answers to All Questions, Analytical Chemistry, Principles and Techniques** Larry G. Hargis, 1987-12      **Catalog of Copyright Entries. Third Series** Library of Congress. Copyright Office, 1959 Includes Part 1 Number 1 Books and Pamphlets Including Serials and Contributions to Periodicals January June

*Critical Care Paramedic* Bryan E. Bledsoe, Randall W. Benner, 2006 Critical Care Paramedic is the first and only book written for the expanding field of critical care transport Utilizing the successful format of Brady's best selling Paramedic Care Principles Practice this book is both written and extensively reviewed by leading experts in the field This comprehensive book addresses all aspects of critical care transport including operations safety advanced pharmacodynamics advanced physiologic monitoring advanced airway management medical emergencies trauma emergencies organ donation and other important areas of critical care Critical Care Paramedic will be the industry standard for critical care paramedic education and available only from Brady

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variations and balancing these reactions providing a material balance between starting compounds and reactions products This book explores that element of the teaching of the fundamentals of chemical literacy writing complete equations of chemical reactions and balancing them It contains 49 figures 22 schemes and 12 tables and 93 problems with answers This book will be very useful for high school students interested in chemical sciences higher education teachers students in colleges and universities majoring in chemistry and biochemistry and chemistry professional working in industry It also contains information about properties of the most common elements and applications of a variety of their chemical compounds

**Essentials of Chemistry** Ralph A. Burns, John William Hill, 1995 The second edition of this chemistry textbook that uses practice examples and applications relating chemistry to our lives and the environment Organic Chemistry Ray Quincy Brewster, William Edwin McEwen, 1959 Physiological Psychology Timothy K. Smock, 1999

Examining changes in the field this work emphasizes neuroscience orientation Humanities themes and examples are used to illuminate complex topics and natural science concepts are taught throughout **Supervision for Quality Education in**

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**Study Guide [to Accompany] General Chemistry** Dixie Goss, 2002 **Science and Technology** , 1963 **Quantum Chemistry** Robert L. Flurry, 1983 **Student's Guide, Chemistry, the Central Science** James C. Hill, 1991 *Basic Chemistry* G. William Daub, William S. Seese, 1996 Written in a style and language that users without science backgrounds can understand This best selling introduction to the basic principles of chemistry draws on the reader's own experiences through analogies and cartoons to learn difficult concepts The clear systematic thinking approach to problem solving has also been highly praised by reviewers and users alike Countdown sections in each chapter consisting of five review questions keyed to previous material provide readers with a basis for material introduced in the new chapter Study exercises found immediately after new topics are introduced reinforce chapter problem material You and Chemistry marginal application icon relates chemistry to the real world End of chapter essays entitled Elements and Compounds relate the applications of specific elements or compounds to the reader's life Study Guide and Partial Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry Susan McMurry, 1996 This internationally acclaimed detective series is just the thing for lovers of those Number One Ladies looking for a darker more realistic view of Botswana Sue Baker Publishing News **A Voyage of Exploration** , 1986

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