

**Double-Replacement Precipitation Lab**

**Introduction:** Double-Replacement reactions occur when two ionic solutions mix to form an insoluble solid. If all cations and anions are soluble then no reaction will occur. During the lab procedure, several ionic solutions will be mixed to determine if a precipitate reaction will occur. Those reactions that do not form a precipitate will be marked for No Reaction (NR). Students will practice writing and balancing double-replacement reactions.

**Pre-Lab questions:**

1. How would you describe a double replacement reaction to a friend who is struggling with reaction types?
2. What is a precipitate?

**Materials:**

- 0.2M sodium bromide, NaBr
- 0.2M sodium carbonate,  $\text{Na}_2\text{CO}_3$
- 0.2M sodium phosphate,  $\text{Na}_3\text{PO}_4$
- 0.2M sodium sulfate,  $\text{Na}_2\text{SO}_4$
- 0.2M sodium hydroxide, NaOH
- Transparency paper
- 0.2M lithium chloride, LiCl
- 0.2M copper(II) chloride,  $\text{CuCl}_2$
- 0.2M zinc chloride,  $\text{ZnCl}_2$
- 0.2M strontium chloride,  $\text{SrCl}_2$
- 0.2M barium chloride  $\text{BaCl}_2$
- Background paper Template

**Safety:** You MUST wear goggles during this lab activity

**Procedure:**

1. Place the overhead sheet over the background paper.
2. Using the Data Table as a guide, mix every possible pair of solutions in a separate square of the background paper. **Use 1 or 2 drops of each solution.** Do NOT contaminate the individual pipets with different solutions.
3. Mix the solutions with a rinsed toothpick if you are uncertain if a precipitate formed. Rinse the toothpick after EACH mixing – be careful of which end of the rod you use so that you do not touch the chemicals or contaminate your samples!
4. Observe each mixture carefully for signs of a precipitate. (Some precipitates are light in color and hard to see so you may want to slide the transparency over the black lab top on occasion.) Record your observations (presence or absence of a precipitate, and other characteristics such as color, texture, etc.) in the Data Table. If nothing happens, write NR for no reaction.
5. Rinse and wipe the transparency sheet into the sink. Dry the sheet.

**Observations:** Write the observations you made from combining these materials in the spaces in the table below. Note any color changes or precipitate (cloudiness/solid). If no change is observed write NR for no reaction.

	NaBr	$\text{Na}_2\text{CO}_3$	$\text{Na}_3\text{PO}_4$	$\text{Na}_2\text{SO}_4$	NaOH
LiCl					
$\text{CuCl}_2$					
$\text{ZnCl}_2$					
$\text{SrCl}_2$					
$\text{BaCl}_2$					

# Prentice Hall Double Replacement Reaction Lab

**N Noddings**



## **Prentice Hall Double Replacement Reaction Lab:**

*Chemistry* Eugene LeMay, Jr., Herbert Beall, Karen M. Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02

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## **Table of Contents Prentice Hall Double Replacement Reaction Lab**

1. Understanding the eBook Prentice Hall Double Replacement Reaction Lab
  - The Rise of Digital Reading Prentice Hall Double Replacement Reaction Lab
  - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Double Replacement Reaction Lab
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Prentice Hall Double Replacement Reaction Lab
  - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Double Replacement Reaction Lab
  - Personalized Recommendations
  - Prentice Hall Double Replacement Reaction Lab User Reviews and Ratings
  - Prentice Hall Double Replacement Reaction Lab and Bestseller Lists
5. Accessing Prentice Hall Double Replacement Reaction Lab Free and Paid eBooks

- Prentice Hall Double Replacement Reaction Lab Public Domain eBooks
  - Prentice Hall Double Replacement Reaction Lab eBook Subscription Services
  - Prentice Hall Double Replacement Reaction Lab Budget-Friendly Options
6. Navigating Prentice Hall Double Replacement Reaction Lab eBook Formats
- ePub, PDF, MOBI, and More
  - Prentice Hall Double Replacement Reaction Lab Compatibility with Devices
  - Prentice Hall Double Replacement Reaction Lab Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Prentice Hall Double Replacement Reaction Lab
  - Highlighting and Note-Taking Prentice Hall Double Replacement Reaction Lab
  - Interactive Elements Prentice Hall Double Replacement Reaction Lab
8. Staying Engaged with Prentice Hall Double Replacement Reaction Lab
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Prentice Hall Double Replacement Reaction Lab
9. Balancing eBooks and Physical Books Prentice Hall Double Replacement Reaction Lab
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Prentice Hall Double Replacement Reaction Lab
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Prentice Hall Double Replacement Reaction Lab
- Setting Reading Goals Prentice Hall Double Replacement Reaction Lab
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Prentice Hall Double Replacement Reaction Lab
- Fact-Checking eBook Content of Prentice Hall Double Replacement Reaction Lab
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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