

## 14

## THE BEHAVIOR OF GASES

## Chapter Test II

## A. Matching

Match each item in Column I with the correct description in Column II. Write the letter of the correct description next to the item.

## Column I

1. The constant relating unit temperature changes to changes in volume for a fixed amount of gas at constant pressure (the ratio of the change in volume to the change in temperature).
2. The volume of a fixed amount of a gas at constant pressure and temperature.
3. The ratio of the mass of a gas to its volume, independent of the actual amount of gas used.
4. An experimental gas constant determined from the ideal gas law at 0°C and 1 atm.
5. A constant of nature that has values of about 8.314 J/mol·K.
6. The a parameter of gas constant temperatures for which the gas molecules move with greatest energy.
7. An experimental gas constant derived from the ideal gas law at 0°C and 1 atm.
8.  $\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$
9. The amount of heat enough to raise 1 g of water 1°C.
10. A unit of energy (joules).

## Column II

- a. pressure-volume work
- b. ideal gas constant  $R$
- c. volume
- d. compressibility
- e. Boyle's law
- f. joules
- g. Boyle's constant
- h. standard temperature and pressure
- i. Celsius
- j. Fahrenheit
- k. heat capacity of oxygen

## B. Multiple Choice

Choose the correct answer and write its letter on the line.

1. Expanding one volume of a monatomic gas from 1 atm and 100°C to 1 atm and 200°C requires \_\_\_\_\_ J of energy.  
a. 100 joules  
b. 1000 joules  
c. 10000 joules  
d. 100000 joules

# Prentice Hall Chemistry 8 Test

**Thomas J. Bruno, Paris D.N. Svoronos**

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