

# 13.1 THE NATURE OF GASES

P.1

## Section Review

### Objectives

- Describe the assumptions of the kinetic theory as it applies to gases
- Interpret gas pressure in terms of kinetic theory
- Define the relationship between Kelvin temperature and average kinetic energy

### Vocabulary

- |                  |                        |                             |
|------------------|------------------------|-----------------------------|
| • kinetic energy | • vacuum               | • pascal (Pa)               |
| • kinetic theory | • atmospheric pressure | • standard atmosphere (atm) |
| • gas pressure   | • barometer            |                             |

### Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

The kinetic theory describes the 1 of particles in matter and the forces of attraction between them. The theory assumes that the volume occupied by a gas is mostly 2, that the particles of gas are relatively 3, move 4 of each other, and are in constant 5 motion. The 6 between particles are perfectly elastic so that the total 7 remains constant. Gas pressure results from the simultaneous collisions of billions of particles with an object. Barometers are used to measure 8 pressure. Standard conditions are defined as a temperature of 9 and a pressure of 10.

- motion
- empty space
- far apart
- independently
- random / rapid
- collisions
- Kinetic energy
- atmospheric
- 0°C / 273 K
- 1 atm / 101.3 kPa

### Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

ST 11. Atmospheric pressure is 760 mm Hg.

AT 12. The SI unit of pressure is the pascal.

↳ a multiple of kilo pascal

18 pts  
 - 2 pt if late  
 - 1 pt if NO ID

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**Scott C. Dulebohn**



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