

CHAPTER 6 REVIEW*Chemical Bonding***MIXED REVIEW**

SHORT ANSWER Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

_____ lattice energy _____	a. ionic bond
_____ bond energy _____	b. covalent bond
_____ enthalpy of vaporization _____	c. metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

_____ ionic _____	a. H—F	_____ nonpolar covalent _____	d. H—H
_____ ionic _____	b. Na—Cl	_____ polar covalent _____	e. H—C
_____ polar covalent _____	c. H—O	_____ polar covalent _____	f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen
atom and the unshared electron pair of a strongly electronegative atom such as O,
N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared,
the hydrogen bond is a weaker attraction. Hydrogen bonds are generally
intermolecular, while ionic and covalent bonds occur between ions or atoms
respectively.

4. H₂S and H₂O have similar structures and their central atoms belong to the same group. Yet H₂S is a gas at room temperature and H₂O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond.
Increased polarity in H₂O bonds means a stronger intermolecular attraction, making
water a liquid at room temperature. Hydrogen bonding exists between water
molecules, but not between hydrogen sulfide molecules.

Prentice Hall Chemistry Chapter 8 Covalent Bonding Answers

William S. Seese, Guido H. Daub



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Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction **Prentice Hall Chemistry**, 2000 *Chemistry* Eugene LeMay, Jr., Herbert Beall, Karen M. Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02 Computational Chemistry Errol G. Lewars, 2010-11-09 This corrected second edition contains new material which includes solvent effects the treatment of singlet diradicals and the fundamentals of computational chemistry Computational Chemistry Introduction to the Theory and Applications of Molecular and Quantum Mechanics is an invaluable tool for teaching and researchers alike The book provides an overview of the field explains the basic underlying theory at a meaningful level that is not beyond beginners and it gives numerous comparisons of different methods with one another and with experiment The following concepts are illustrated and their possibilities and limitations are given potential energy surfaces simple and extended Hueckel methods ab initio AM1 and related semiempirical methods density functional theory DFT Topics are placed in a historical context adding interest to them and removing much of their apparently arbitrary aspect The large number of references to all significant topics mentioned should make this book useful not only to undergraduates but also to graduate students and academic and industrial researchers **Prentice Hall Biology** Sandra Gottfried, Gerry M. Madrazo, Jr., 1990 Physical Chemistry Robert J. Silbey, Robert A. Alberty, George A. Papadantonakis, Mounqi G. Bawendi, 2022-06-15 Ever since Physical Chemistry was first published in 1913 it has remained a highly effective and relevant learning tool thanks to the efforts of physical chemists from all over the world Each new edition has benefited from their suggestions and expert advice The result of this remarkable tradition is now in your hands *Chemistry* James C. Hill, 2003 This book assists students through the text material with chapter overviews learning objectives review of key terms cumulative chapter review quizzes and self tests Included are answers to all Student Guide exercises Chapter summaries are correlated to those in the Instructor's Resource Manual *Student's Guide, Chemistry, the Central Science* James C. Hill, 1991 *Chemistry* Charles H. Corwin, 1994 The book focuses on the concepts of chemistry and the applications that maintain and generate motivation for the subject of chemistry *Chemistry* Theodore Lawrence Brown, 1988 **Chemistry - The Central Science** James C. Hill, Bruce Edward Bursten, 2006 Chemistry The Central Science is the most trusted book on the market its scientific accuracy clarity innovative pedagogy functional problem solving and visuals set this book apart Brown LeMay and Bursten teach students the concepts and skills they need without overcomplicating the subject A comprehensive media package that works in tandem with the text helps students practice and learn while providing instructors the tools they need to succeed Publisher's

description *Collected Accounts of Transition Metal Chemistry* Fred Basolo, Joseph F. Bunnett, Jack Halpern, 1973

Physical Chemistry John S. Winn, 1995 This text presents physical chemistry as a coherent whole rather than a set of disjointed topics and shows how the subject relates to the rest of chemistry and physics It emphasizes physical models as well as mathematical techniques along with both rigorous and approximate order of magnitude problem solving Designed to progress beyond a numerical answer problems expose the physical significance of the situation and teach students how to pose a problem in the first place In addition modern molecular concepts currently unanswered problems in research experimental techniques and new directions in the field are introduced wherever appropriate An orderly progression of thermodynamics carefully builds students knowledge without covering too much too early on Chemical reaction thermodynamics is covered in Chapter 7 after the culmination of thermodynamics with advanced material in Chapter 10

Fundamentals of Chemistry Ralph A. Burns, 1995 Basic Chemistry William S. Seese, Guido H. Daub, 1985

Chemical Matter Prentice-Hall Staff, 1994 Authorized teaching resource in Alberta for senior high science 14 24 1995 2004 **Technical Book Review** ,1963 **Study Guide and Solutions Manual, Fundamentals of General, Organic,**

and Biological Chemistry, Third Edition John McMurry, Susan McMurry, 1999 Provides worked out solutions to text problems along with chapter by chapter outlines and a variety of self tests at the end of each chapter **Clear and Simple**

Chemistry James Berger, Jim Berger, 1986 **Shriver and Atkins' Inorganic Chemistry** Peter Atkins, 2010 Inorganic Chemistry fifth edition represents an integral part of a student s chemistry education Basic chemical principles are set out clearly in Foundations and are fully developed throughout the text culminating in the cutting edge research topics of the Frontiers which illustrate the dynamic nature of inorganic chemistry

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