

## CHEMISTRY: LIMITING REAGENT Practice Problems

Please show your work on a separate piece of paper!

Assume standard temperature and pressure, STP conditions for all reactions.

1. At high temperatures, sulfur combines with iron to form the brown-black iron (II) sulfide:



In one experiment, 7.62 g of Fe are allowed to react with 8.67 g of S.

- What is the limiting reagent, and what is the reactant in excess?
- Calculate the mass of FeS formed.

2. Acrylonitrile,  $\text{C}_3\text{H}_3\text{N}$ , is the starting material for the production of a kind of synthetic fiber (acrylics) and can be made from propylene,  $\text{C}_3\text{H}_6$ , by reaction with nitric oxide,  $\text{NO}$ , as follows:



What mass of  $\text{C}_3\text{H}_3\text{N}$  can be made when 21.6 g of  $\text{C}_3\text{H}_6$  react with 21.6 g of nitric oxide?

3. Calculate the percent yield for the reaction:  $\text{P}_{4(s)} + 6 \text{Cl}_{2(g)} \rightarrow 4 \text{PCl}_{3(l)}$   
if 75.0 g of phosphorus reacts with excess chlorine gas to produce 111.0 g of phosphorus trichloride.

4. Formic acid,  $\text{HCHO}_2$ , burns in oxygen to form carbon dioxide and water as follows:



If a 3.15-g sample of formic acid was burned in 2.0 L of oxygen, what volume of carbon dioxide would be produced?

5. Zinc metal reacts with hydrochloric acid to produce zinc chloride and hydrogen gas.
- Balance the following reaction:  $\text{Zn}_{(s)} + \text{HCl}_{(aq)} \rightarrow \text{ZnCl}_{2(aq)} + \text{H}_{2(g)}$
  - A 3.50-g sample of zinc metal is allowed to react with 2.50 g of hydrochloric acid.

Complete the following table:

reactants/products	Zn (grams)	HCl (grams)	ZnCl <sub>2</sub> (grams)	H <sub>2</sub> (L)
Before reaction				
After reaction	1.26 g			

# Prentice Hall Chemistry Practice Problems Limiting Reactant

**James C. Hill**

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular gradient shape partially visible behind it.

## **Prentice Hall Chemistry Practice Problems Limiting Reactant:**

**Prentice Hall Chemistry**, 2000      Problems and Problem Solving in Chemistry Education Georgios

Tsaparlis, 2021-05-17 Problem solving is central to the teaching and learning of chemistry at secondary tertiary and post tertiary levels of education opening to students and professional chemists alike a whole new world for analysing data looking for patterns and making deductions As an important higher order thinking skill problem solving also constitutes a major research field in science education Relevant education research is an ongoing process with recent developments occurring not only in the area of quantitative computational problems but also in qualitative problem solving The following situations are considered some general others with a focus on specific areas of chemistry quantitative problems qualitative reasoning metacognition and resource activation deconstructing the problem solving process an overview of the working memory hypothesis reasoning with the electron pushing formalism scaffolding organic synthesis skills spectroscopy for structural characterization in organic chemistry enzyme kinetics problem solving in the academic chemistry laboratory chemistry problem solving in context team based active learning technology for molecular representations IR spectra simulation and computational quantum chemistry tools The book concludes with methodological and epistemological issues in problem solving research and other perspectives in problem solving in chemistry With a foreword by George Bodner

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*Chemistry* Charles H. Corwin, 1994 The book focuses on the concepts of chemistry and the applications that maintain and generate motivation for the subject of chemistry **Student's Guide, Chemistry, the Central Science** James C. Hill, 1991

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*Modelling and Numerical Simulations II* Mordechai Schlesinger, 2009-08-15 The present volume is the second in a two volume set dealing with modelling and numerical simulations in electrochemistry Emphasis is placed on the aspect of nanoelectrochemical issues It seems appropriate at this juncture to mention the growing body of opinion in some circles that George Box was right when he stated three decades ago that All models are wrong but some are useful Actually when the statement itself was made it would have been more appropriate to say that All models are inaccurate but most are useful nonetheless At present however the statement as it was made is far more appropriate and closer to the facts than ever before Currently we are in the midst of the age of massively abundant data Today's philosophy seems to be that we do not need to know why one piece of information is better than another except through the statistics of incoming and outgoing links between information and this is good enough It is why both in principle and in practice one can translate between two languages without knowledge of either While none of this can be ignored and it may even be true that All models are wrong and increasingly you can succeed without them the traditional approach of scientific modelling is still the order of the day That approach may be stated as hypothesize measure model test It is in this light that the present volume should be viewed **Chemistry** Theodore Lawrence Brown, 1988 Study Guide to

Accompany Elements of Chemistry, General, Organic, and Biological -- Robert Boikess/Kenneth Breslauer/Edward Edelson Robert A. Klein, 1986 **Atoms, Molecules, and Reactions** Ronald James Gillespie, 1994 Traditional college level chemistry including principles

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