

## Chemistry - Review Chapters (415) 6

Directions: Answer each of the following questions. You need not use complete sentences.

1. By what property did Mendeleev arrange the elements?

**ATOMIC MASS**

2. By what property did Moseley suggest that the periodic table be arranged?

**ATOMIC NUMBER**

3. What is a period? How many are there in the periodic table?

**A HORIZONTAL ROW IN THE PERIODIC TABLE; 7**

4. What is a group (also called a family)? How many are there in the periodic table?

**A VERTICAL COLUMN IN THE PERIODIC TABLE; 18**

5. State the number of valence electrons in an atom of:

a. sulfur	b. calcium	c. chlorine	d. arsenic
<b>6</b>	<b>2</b>	<b>7</b>	<b>5</b>

6. Give the names and chemical symbols for the elements that correspond to these atomic numbers:

a. 10	b. 18	c. 36	d. 90
<b>Ne, NEON</b>	<b>Ar, ARGON</b>	<b>Kr, KRYPTON</b>	<b>Th, THORIUM</b>

7. List, by number, both the period and group of each of these elements.

	<u>Symbol</u>	<u>Period</u>	<u>Group</u>
a. beryllium	Be	<b>2</b>	<b>2</b>
b. iron	Fe	<b>4</b>	<b>8</b>
c. lead	Pb	<b>6</b>	<b>14</b>

8. Which of the following pairs of elements belong to the same period?

<b>a. Na and Cl</b>	b. Na and Li	c. Na and Cu	d. Na and Ne
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9. Which of the following pairs of elements belong to the same group?

a. H and He	b. Li and Be	<b>c. C and Pb</b>	d. Ga and Ge
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10. How does an element's period number relate to the number of the energy level of its valence electrons?

**PERIOD NUMBER = ENERGY LEVEL OF VALENCE ELECTRONS**

11. What are the transition elements?

**GROUPS 3-12**

12. In what type of orbitals are the actinide and lanthanide electrons found?

**f ORBITALS**

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