

Component code numbers

Codes d'appareils

Números de código de componentes

TV and VCR
Téléviseurs et magnétoscopes
Televisores y videocaseteras

Brand Marque Marca	Component Appareil Componente	Code No. Code Num. de código
Sony	VCR	00 (VTR) (Base) 02 (VTR) 8 mm 03 (VTR) 10mm 04, 05, 06 (VTR) 1.5 in. DV
	TV	01
	VCR	02
	TV	03, 04, 05
Acorn	VCR	41, 50, 44
Alai	VCR	51, 54
AOC	TV	12
Audio Dynamic	VCR	17, 09
B&W (Hawaii)	VCR	21, 46
	TV	34, 08, 38
Bosconi	VCR	35, 21, 07
	TV	04, 12
Carm	VCR	11, 01
Ciccone	VCR	03
Craig	VCR	04, 05
	TV	13, 14, 06, 10
Coleman	VCR	18
Cosmos	TV	11
Curtis Mathis	VCR	07, 01, 02
	TV	05, 04, 23
Daphne	VCR	44, 14, 02
	TV	15, 06, 03, 40
	TV/VCR & Combo	04 & 44
DBS	VCR	17, 36, 40
Diagrams	VCR	07
Electron	VCR	44, 23, 25, 04, 26, 21, 41
	TV	21, 26, 25, 23, 05, 14, 40
	TV/VCR & Combo	21 & 21, 21
Fuller	VCR	11, 13, 36, 38
	TV	27
Fujitsu	VCR	41

Brand Marque Marca	Component Appareil Componente	Code No. Code N°m. de código
General Electric	VCR	31, 37, 39
	TV	33, 38, 39
	TV/VCR/Combo	37 & 38, 39
Columbia	VCR	39
	TV	33, 33, 38, 39
	TV/VCR/Combo	TV/Combo & 38
De Vries	VCR	31, 43, 49
Hitachi	VCR	31, 37, 38, 39
	TV	31, 33, 32
IBC	VCR	39
Instant Replay	VCR	31, 33
K' Pong	VCR	31, 38, 39, 39, 39, 39, 39, 39
	TV	33, 38, 39
	VCR	31, 47, 48, 49, 39, 49
	TV	38
Kennwood	VCR	31, 38, 39, 49
KTC	TV	31, 33, 33, 38
L&L/Seaco	VCR	41, 39, 38, 38, 37, 33, 38
	TV	33, 38, 37, 39, 38, 37
Magnavox	VCR	31, 39, 32, 39
	TV	31, 33, 33, 43, 47
	TV/VCR/Combo	31, 38, 47 & 32, 33
Matsushita	VCR	31, 38, 49
Matsui	VCR	34
Mitsumi	VCR	31, 38
Mosconi	VCR	38, 37
Multimedia/	VCR	38, 37, 38, 39, 38, 37
NACA	TV	34, 33, 32
Multitech	VCR	38, 43, 38
	VCR	31, 38, 49
	TV	33, 33
	VCR	31, 33
Optima	VCR	37
Orion	VCR	39
Panasonic	VCR	38, 33, 33, 33
	TV	38, 38, 38
	TV/VCR/Combo	34 & 33, 33
Pioneer	VCR	38, 37

Brand/Make/Model	Component/Applied Component	Code No./Code Num./de código
Pioneer	TC	11, 12
	TR	13, 15
Philips	VCR	13, 14, 15
	TR	14
Pioneer	VCR	14
	TR	15, 16, 19
Portland	TR	15
Quasar	VCR	14, 15, 16, 19, 20
	TR	16
	TR/VCR Combo	15 & 16, 19, 20
Radio Shack	TR	15, 17, 18, 23
RAI/A/	VCR	17, 18, 19, 21, 22, 24, 25
PHILCO/A/N		16, 20
	TR	15, 16, 18, 19, 24, 25
	TR/VCR Combo	15 & 16
Realistic	VCR	14, 15, 16, 18, 19, 20
Samson	VCR	17
Sanyo	VCR	16
Sanyo	VCR	25, 26, 28
	TR	26, 27, 29
	TR/VCR Combo	15, 17 & 24, 25, 26
Sanyo	VCR	19, 20
	TR	25
Seash	VCR	15, 16, 24, 26, 26, 27, 28, 29
	TR	25, 27
Sharp	VCR	20, 21
	TR	16, 23
Shionon	VCR	18
Signature 2000	VCR	43, 55, 56
	TR	44, 58
Sylvania	VCR	15, 17, 22, 41
	TR	17, 23, 57
Symphonix	VCR	41
	TR	23, 58, 60
	TR/VCR Combo	15 & 41
Tekono	VCR	45
Tosong	VCR	17, 59, 60
Toson	VCR	17, 58, 61, 62
Toshiba	VCR	12, 13
Toshiba	VCR	15
	TR	16, 18
Toshiba	TR	19, 22, 23
Wards	VCR	50, 51, 52, 54, 55
	TR	45, 53, 52
Yamaha	VCR	23, 17, 38, 40
Zenith	VCR	16
	TR	19, 20

Cable Box
Decodereurs
Adaptadores para cablevisión

Brand/Markas/Marka	Code No./Kode No./Kode No. de código
Blonde/Regel	01, 02, 04, 06
Blonde/PC 1	01, 02, 03, 04, 05, 06, 07, 08, 09, 10
Club	06, 09
Panasonic	01
Prestige	04, 05
Scotch/Scotch Atlantic	04, 11
Tecumseh	04
Vermeer	02, 10

DSS Tuner
Récepteurs DSS
Sintonizadores del sistema de
radiodifusión digital

Brand/Marken/Marca	Estim. No. Queda/Noim. de colágeno
Isomax	100
Elmex	100
Eliglo	100
Insight	100
Paracore	100
Br ¹ , A ² , P ³ (Br ¹ , A ²)	100, 100
Fundalga	100

Sony Rm D2 Universal Remotes Owners Manual

Xiaolong Qi



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Table of Contents Sony Rm D2 Universal Remotes Owners Manual

1. Understanding the eBook Sony Rm D2 Universal Remotes Owners Manual
 - The Rise of Digital Reading Sony Rm D2 Universal Remotes Owners Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Sony Rm D2 Universal Remotes Owners Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Sony Rm D2 Universal Remotes Owners Manual
 - User-Friendly Interface

4. Exploring eBook Recommendations from Sony Rm D2 Universal Remotes Owners Manual
 - Personalized Recommendations
 - Sony Rm D2 Universal Remotes Owners Manual User Reviews and Ratings
 - Sony Rm D2 Universal Remotes Owners Manual and Bestseller Lists
5. Accessing Sony Rm D2 Universal Remotes Owners Manual Free and Paid eBooks
 - Sony Rm D2 Universal Remotes Owners Manual Public Domain eBooks
 - Sony Rm D2 Universal Remotes Owners Manual eBook Subscription Services
 - Sony Rm D2 Universal Remotes Owners Manual Budget-Friendly Options
6. Navigating Sony Rm D2 Universal Remotes Owners Manual eBook Formats
 - ePub, PDF, MOBI, and More
 - Sony Rm D2 Universal Remotes Owners Manual Compatibility with Devices
 - Sony Rm D2 Universal Remotes Owners Manual Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Sony Rm D2 Universal Remotes Owners Manual
 - Highlighting and Note-Taking Sony Rm D2 Universal Remotes Owners Manual
 - Interactive Elements Sony Rm D2 Universal Remotes Owners Manual
8. Staying Engaged with Sony Rm D2 Universal Remotes Owners Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Sony Rm D2 Universal Remotes Owners Manual
9. Balancing eBooks and Physical Books Sony Rm D2 Universal Remotes Owners Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Sony Rm D2 Universal Remotes Owners Manual
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Sony Rm D2 Universal Remotes Owners Manual
 - Setting Reading Goals Sony Rm D2 Universal Remotes Owners Manual
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Sony Rm D2 Universal Remotes Owners Manual
 - Fact-Checking eBook Content of Sony Rm D2 Universal Remotes Owners Manual
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
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web 1 consider the reaction $\text{I}_2\text{O}_5 + 5\text{CO} \rightarrow 5\text{CO}_2 + \text{I}_2$ a 80.0 grams of iodine pentoxide I_2O_5 reacts with 28.0 grams of carbon monoxide CO determine the mass of iodine I_2 which could be produced b if in the above situation only 0.160 moles of iodine I_2 was produced i what mass of iodine was produced

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web limiting reactant theoretical and percentage yields key $4\text{KO}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{KOH} + 3\text{O}_2$ g if a reaction vessel contains 0 mol KO_2 and 0 mol H_2O what is the limiting reactant

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web limiting reactant reactant present in an amount lower than required by the reaction stoichiometry thus limiting the

amount of product generated percent yield measure of the efficiency of a reaction expressed as a percentage of the theoretical yield

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation

$$8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$$
 is your theoretical yield so percent yield

$$\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$$
 problems 6 & 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH_3 can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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web limiting reagents problem one 1 for the balanced equation

$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$
 if the reaction of 21.3 grams of O_2 produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have

$$21.3 \text{ g } \text{O}_2 \times \frac{1 \text{ mole } \text{O}_2}{32 \text{ g } \text{O}_2} = 0.666 \text{ moles } \text{O}_2$$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency mass of product obtained / calculated mass of the product expected $\times 100$ in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction

$$2\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$$
 b determine the theoretical yield of Al_2O_3

$$160.0 \text{ g } \text{Al} \times \frac{1 \text{ mole } \text{Al}}{27 \text{ g } \text{Al}} \times \frac{2 \text{ mole } \text{Al}_2\text{O}_3}{4 \text{ mole } \text{Al}} \times 101.96 \text{ g } \text{Al}_2\text{O}_3 = 302.3 \text{ g } \text{Al}_2\text{O}_3$$
 c determine the percent yield

$$\frac{260.0 \text{ g } \text{Al}_2\text{O}_3}{302.3 \text{ g } \text{Al}_2\text{O}_3} \times 100 = 86.01\%$$
 theoretical 302.3 g

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web 7 Na₂C₂O₄ is the limiting reactant percent yield 86.68 only four molecules can be made 9 this amount cannot be

weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of h₂ are left over 13

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web the percent yield is calculated as follows text percent yield $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent yield theoretical yield based on this definition we would expect a percent yield

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