

# Physical Science – 2nd Semester – Final Exam Study Guide – (2012–2013)

## 1. Physical Science Basics

- What tool(s) would you want to use to find the:
  - Mass of an object? triple beam balance Basic SI Unit? kg
  - Volume of an object? ruler or water displacement method Basic SI Unit? cm<sup>3</sup> or ml
  - Length of an object? Ruler Basic SI Unit? meter
  - Density of an object? graduated cylinder Basic SI Unit? ml
- What is the formula for calculating density? mass/volume
  - What is the density of a cube that is 25 g and has a volume of 3 cm<sup>3</sup>? 8g/cm<sup>3</sup>
  - Compare the density of solid H<sub>2</sub>O to liquid H<sub>2</sub>O, and explain what observable effect this has. solid H<sub>2</sub>O less dense than liquid H<sub>2</sub>O

## 2. Chemistry Basics

- Element (def.) substance made up of one kind of atom
- Atom (def.) basic building block of matter
- Compound (def.) a molecule that contains atoms or molecules of two or more different elements
- Molecule (def.) a combination of two or more of the same or different atoms that are bonded together

## 3. Model of the Atomic Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created models to describe them.

Why have atomic theories changed over time? Better technology came around.

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	1808	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	Billiard Ball Model
J.J. Thomson	1897	Negatively charged electron	Plum Pudding Model
Niels Bohr	1913	Electrons move <u>fixed orbits</u> around the atom's nucleus in specific layers, or <u>electron shells</u> . Atoms absorb or give off energy when the electrons move from one shell ( <u>energy level</u> ) to another.	Bohr Model
Ernest Rutherford	1909	1909 /Based on idea that positive reflects positive/ atom mostly empty space; electrons orbit randomly around positively charged nucleus	Planetary Model

## 4. Periodic Table

- The current periodic table is arranged by increasing atomic number.
- Reading the Periodic Table
  - An element's properties can be predicted from its location in the periodic table.
  - The atomic number represents the number of protons.
  - The atomic mass represents the number of protons and neutrons.

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**Sharon Wynne**



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*PLACE Science 05 Teacher Certification Test Prep Study Guide* Sharon Wynne, 2007-05. This guide aligned specifically to standards prescribed by the Colorado Department of Education. Includes 21 competencies/skills found on the PLACE Science test and 125 sample test questions on scientific inquiry and connections, physical science, life science, and Earth and space science.

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In a digital era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Place Science Exam Study Guide**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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