

## Welcome

Congratulations on purchasing your new Plantronics product. This Quick Start Guide contains instructions for setting up and using your Savi W440 (2100 + W400 Headset) Wireless Headset System.

Please refer to the separate **safety instructions** for important product safety information prior to installation or use of the product.

## System Requirements

The minimum system requirements for Savi W440 are as follows:

- Supported Operating Systems: Windows XP®; Windows Vista®; Windows® 7
- Processor: 400 MHz Pentium® processor or equivalent (Minimum); 1 GHz Pentium processor or equivalent (Recommended)
- RAM: 512 MB (Minimum); 2GB MB (Recommended)
- Hard Disk: 200 MB (Minimum); 500 MB (Recommended)
- Display: 800 x 600, 256 colors (Minimum); 1024 x 768 high color, 32-bit (Recommended)
- Internet Explorer® v6 or better must be on the user's system (Required)

## Complete User Guide and Plantronics Software

We highly recommend that you refer to the complete user guide located on the CD. The user guide provides full details on how to use your system.

We also recommend that you install the Plantronics Unified Runtime Engine and Control Panel located on the CD. With Plantronics software installed you can:

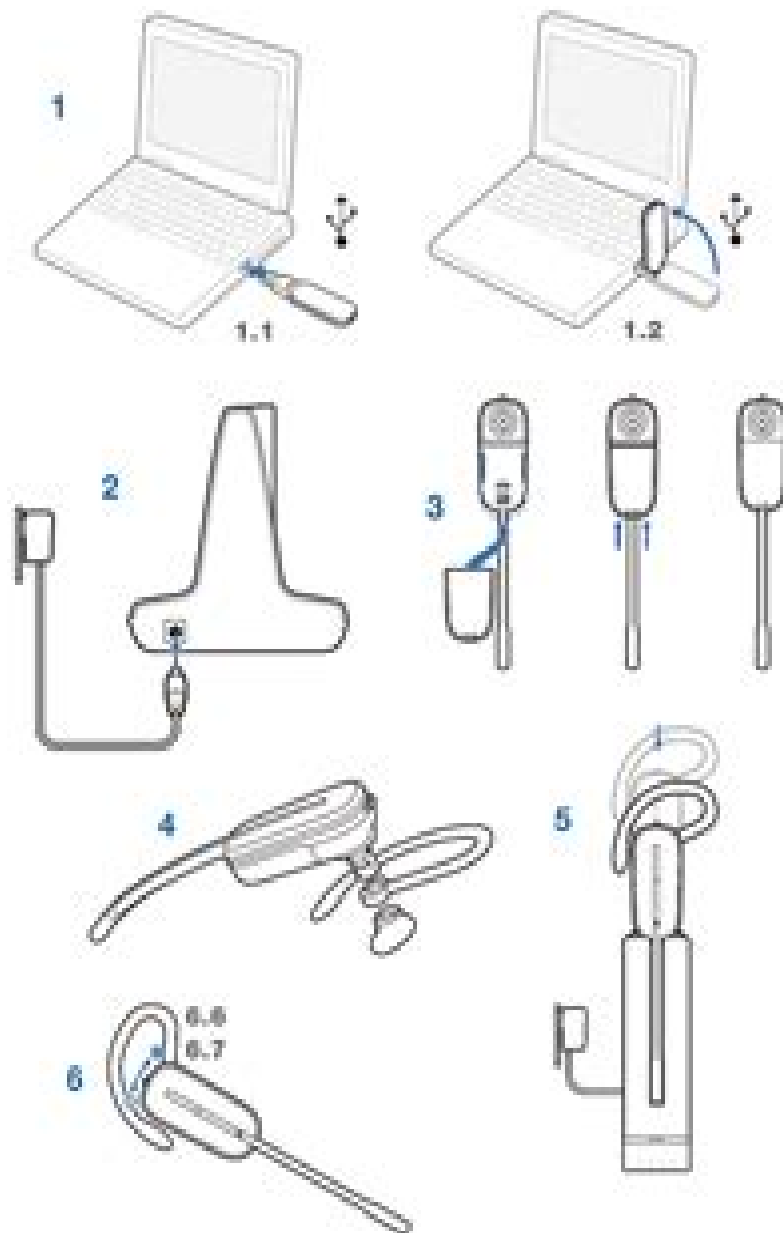
- Remotely answer incoming softphone calls
- Customize your headset and USB adapter settings to your personal preferences
- Automatically manage your PC multimedia so you never miss a call
- Quickly access links to product support and software/firmware upgrade information

## Registering Your Product

Visit [plantronics.com/productregistration](http://plantronics.com/productregistration) to register your product online so we can provide you with the best service and technical support.

## Getting Assistance

The Plantronics Technical Assistance Center is ready to assist you. You can find answers to frequently asked questions, ask a question using e-mail, receive service over the Internet, or speak directly with a representative. Visit [plantronics.com/support](http://plantronics.com/support).



# Plantronics Instruction Manual

**United States. Federal  
Communications Commission**



## **Plantronics Instruction Manual:**

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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction  $2\text{Ag}^+ + \text{CrO}_4^{2-} \rightarrow \text{Ag}_2\text{CrO}_4$

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web jan 29 2020 the ionic form of this reaction is  $2\text{Na}^+ + 2\text{Cl}^- + \text{Cu}^{2+} + \text{SO}_4^{2-} \rightarrow 2\text{Na}^+ + \text{SO}_4^{2-} + \text{CuCl}_2$  s the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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web mar 11 2014 solution ionic equation  $\text{H}^+ + \text{Cl}^- + \text{Na}^+ + \text{HCO}_3^- \rightarrow \text{H}_2\text{O} + \text{CO}_2$  g the spectator ions are cl and na net ionic equation  $\text{H}^+ + \text{Cl}^- + \text{Na}^+ + \text{HCO}_3^- \rightarrow \text{H}_2\text{O} + \text{CO}_2$  g h aq hco<sub>3</sub> aq h<sub>2</sub>o l co<sub>2</sub> g here is a video on spectator ions *chapter 8 04 ionic equations chemistry libretexts* - Mar 06 2023

web the six no<sub>3</sub> aq ions and the six na aq ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation  $3\text{Ba}^{2+} + 2\text{PO}_4^{3-} \rightarrow \text{Ba}_3(\text{PO}_4)_2$  s

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web these spectator ions ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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 $\text{AgNO}_3(\text{aq}) + \text{KCl}(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{KNO}_3(\text{aq})$  total ionic  $\text{Ag}^+(\text{aq}) + \text{NO}_3^-(\text{aq}) + \text{K}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{K}^+(\text{aq}) + \text{NO}_3^-(\text{aq})$  net ionic  $\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s})$

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water  $\text{H}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$  the spectator ions in  $\text{KOH}(\text{aq})$  and  $\text{HNO}_3(\text{aq})$  are  $\text{K}^+(\text{aq})$  and  $\text{NO}_3^-(\text{aq})$  how to identify and list the

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web 18 terms lourdes l22 preview terms in this set 4 molecular equation showing the complete chemical formulas of reactions and products  $\text{Pb}(\text{NO}_3)_2(\text{aq}) + 2\text{KCl}(\text{aq}) \rightarrow \text{PbCl}_2(\text{s}) + 2\text{KNO}_3(\text{aq})$  complete ionic equation an equation written with all soluble strong electrolytes shown as ions

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web the  $\text{Ag}^+(\text{aq})$  and  $\text{Cl}^-(\text{aq})$  ions become  $\text{AgCl}(\text{s})$  but the  $\text{Na}^+(\text{aq})$  ions and the  $\text{NO}_3^-(\text{aq})$  ions stay as  $\text{Na}^+(\text{aq})$  ions and  $\text{NO}_3^-(\text{aq})$  ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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web aug 8 2022 a spectator ion is an ion that does not take part in the chemical reaction and is found in solution both

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web the new format requires students to write balanced chemical equations showing only the reacting substances excluding spectator ions for three different sets of reactants and to answer a short question requiring no calculator about each reaction the questions are intended to test the students understanding of the meaning of the reactions

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web jun 2 2023 the spectator ions are cut out from a reaction hence the net ionic equation for the above reaction would be  $2\text{Cl}^- \text{aq} + \text{Cu}^{2+} \text{aq} \rightarrow \text{Cu}^{2+} \text{aq} + 2\text{Cl}^- \text{aq}$  these spectator ions are not actively involved in the reaction the existence of spectator ions affects the Debye length

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