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web may 30 2014 answer link nuclear equations can be solved quite simply let s do one example of alpha decay and one of beta decay first a quick revision of radioactive decay during alpha decay an alpha particle is emitted from the nucleus it is the equivalent of a helium atom i e it has a mass of 4 and an atomic number of 2

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web nuclear reactions can be represented by balanced equations of nuclei in the ${}^A_Z\text{X}$ form ${}^A_Z\text{X}$ notation for atomic nuclei the top number A represents the nucleon number or the mass number nucleon number a total number of protons and neutrons in the nucleus the lower number Z represents the proton or atomic number

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web to describe a nuclear reaction we use an equation that identifies the nuclides involved in the reaction their mass numbers and atomic numbers and the other particles involved in the reaction types of particles in nuclear reactions many entities can be involved in nuclear reactions

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web the carbon 14 nucleus is unstable it s radioactive it s going to undergo spontaneous decay it s going to try to get a better neutron to proton ratio so let s look at the nuclear equation which represents the spontaneous decay of carbon 14 so here is

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web to describe a nuclear reaction we use an equation that identifies the nuclides involved in the reaction their mass numbers and atomic numbers and the other particles involved in the reaction types of particles in nuclear reactions

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web in alpha decay an alpha particle is ejected from an unstable nucleus so here's our unstable nucleus uranium 238 an alpha particle has the same composition as a helium nucleus we saw the helium nucleus in the previous video there are two protons in the helium nucleus and two neutrons

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web sep 12 2022 the half life of the nuclei and thus the material is $t_{1/2} = 693$ t $t_{1/2} = 693$ million years 35 a the decay constant is $\lambda = 1.99 \times 10^{-5} \text{ s}^{-1}$ $\lambda = 1.99 \times 10^{-5} \text{ s}^{-1}$ b since strontium 91 has an atomic mass of 90.90 g

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web complete the nuclear equation to find the atomic number of californium $^{242}_{96}\text{Cm} \rightarrow ^{24}_{10}\text{Ne} + ^{218}_{86}\text{Cf}$ by firing heavier particles such as carbon or boron ions at the target materials heavier elements can be synthesised

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web nuclear equations represent the reactants and products in radioactive decay nuclear fission or nuclear fusion instead of chemical equations where it shows the different number of elements is conserved in a reaction in a nuclear reaction the atomic mass and proton number are conserved

