

**Syllabus reference:**

*"Write ionic half-equations and use them to construct full ionic equations."*

**Aim of this worksheet:**

To introduce you to the idea of balancing redox equations by adjusting coefficients and by adding  $\text{OH}^-$ ,  $\text{H}_2\text{O}$ , and  $\text{H}^+$  where necessary

**Task: Balance the following redox reactions:**

- $\text{SO}_2 (\text{g}) + \text{HNO}_2 (\text{aq}) \rightarrow \text{H}_2\text{SO}_4 (\text{aq}) + \text{NO} (\text{g})$
- $\text{Al} (\text{s}) + \text{H}_2\text{SO}_4 (\text{aq}) \rightarrow \text{Al}_2(\text{SO}_4)_3 (\text{aq}) + \text{H}_2 (\text{g})$
- $\text{Au}^{3+} (\text{aq}) + \text{I}^- (\text{aq}) \rightarrow \text{Au} (\text{s}) + \text{I}_2 (\text{s})$
- $\text{S}^{2-} (\text{aq}) + \text{I}_2 (\text{s}) \rightarrow \text{SO}_4^{2-} (\text{aq}) + \text{I}^- (\text{aq})$
- $\text{Br}_2 (\text{aq}) + \text{OH}^- (\text{aq}) \rightarrow \text{Br}^- (\text{aq}) + \text{BrO}_3^- (\text{aq})$
- $\text{H}_2\text{O}_2 (\text{aq}) + \text{ClO}_4^- (\text{aq}) \rightarrow \text{O}_2 (\text{g}) + \text{ClO}_2^- (\text{aq})$
- $\text{Mn} (\text{s}) + \text{HNO}_3 (\text{aq}) \rightarrow \text{Mn}^{2+} (\text{aq}) + \text{NO}_2 (\text{g})$
- $\text{I}_2 (\text{s}) + \text{OCl}^- (\text{aq}) \rightarrow \text{IO}_3^- (\text{aq}) + \text{Cl}^- (\text{aq})$
- $\text{Cr}_2\text{O}_7^{2-} (\text{aq}) + \text{HNO}_2 (\text{aq}) \rightarrow \text{Cr}^{3+} (\text{aq}) + \text{NO}_3^- (\text{aq})$
- $\text{CrO}_4^{2-} (\text{aq}) + \text{S}^{2-} (\text{aq}) \rightarrow \text{Cr}(\text{OH})_3 (\text{s}) + \text{S} (\text{s})$

# Redox Equations Worksheet With Answers

**F Rizvi**



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