

Chemistry Worksheets Class 11 on Chapter 8 Redox Reactions - Set 4

Q-1: Which of the following is not a redox reaction?

- a) $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
- b) $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
- c) Photosynthesis
- d) Cell respiration

Q-2: Which of the following is not correct for an electrochemical cell?

- a) Convert electrical energy to chemical energy
- b) Anode is the negative terminal
- c) Anode is the positive terminal
- d) Cathode is the positive terminal

Q-3: Which reaction involves neither oxidation nor reduction?

- a) $\text{CrO}_4^{2-} \rightarrow \text{Cr}_2\text{O}_7^{2-}$
- b) $\text{Cr} \rightarrow \text{CrCl}_3$
- c) $\text{KMnO}_4 \rightarrow \text{MnO}_2$
- d) $\text{Fe} \rightarrow \text{Fe}_2\text{O}_3 \cdot x\text{H}_2\text{O}$

Q-4: Phosphorus has the oxidation state of +3 in

- a) Orthophosphoric acid
- b) Pyrophosphoric acid
- c) Metaphosphoric acid
- d) Phosphorous acid

Q-5: If a half cell reaction, $\text{C} + \text{e}^- \rightarrow \text{C}^-$ has a large negative reduction potential, it follows that

- a) C is easily oxidised
- b) C is easily reduced
- c) C^- is readily oxidised
- d) C^- is readily reduced

Q-6: What is a redox couple?

Q-7: Identify the species undergoing oxidation and reduction in the following reactions.

- (i) $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
- (ii) $2\text{H}_2\text{S} + \text{O}_2 \rightarrow 2\text{S} + 2\text{H}_2\text{O}$
- (iii) $\text{CH}_2=\text{CH}_2 + \text{H}_2 \rightarrow \text{CH}_3-\text{CH}_3$

Redox Reactions Worksheet 11 1 Answers

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Redox Reactions Worksheet 11 1 Answers:

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Ekborg, Anders Berg, 2010

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