Balancing Redox Reactions Practice Problems

Choose the redox reaction (click on the equation):

In Acida

Br * MnO,
$$\rightarrow$$
 Br, * Mn²⁺

MnO, * H₂C₂O, \rightarrow CO, * Mn²⁺

S₂O₃ * + I₃ \rightarrow I * S₄O₆

As₂O₃ * NO₃ \rightarrow NO * H₃AsO,

Cr₂O₇ * CI \rightarrow Cr²⁺ + CI,

MnO, * SO₃ \rightarrow SO, * MnO, *

In Bose:

$$Cl_2 + Cr(OH)_3 \rightarrow Cl_1 + CrO_4^2$$
 $As_2O_3 + NO_3 \rightarrow NO + H_3AsO_4$
 $Cr_2O_7^2 + Cl_1 \rightarrow Cr_3^3 + Cl_2$
 $OCl_1 + Cr(OH)_4 \rightarrow Cl_1 + CrO_4^3$
 $MnO_4 + SO_3^2 \rightarrow SO_4^2 + MnO_4^2$

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