Kedox Exercise (Part 2)

#### Chapter 20 Worksheet: Redox ANSWERS

 Determine what is oxidized and what is reduced in each reaction. Identify the oxidizing agent and the reducing agent, also.

2Sr + O<sub>2</sub> → 2SrO Sr<sup>0</sup> to Sr<sup>2+</sup>; oxidized/reducing agent O<sup>0</sup> to O<sup>2</sup>; reduced/ox. ag.

2. 2Li + S → Li<sub>2</sub>S Li<sup>0</sup> to Li<sup>1\*</sup>; oxidized/red. ag. S<sup>0</sup> to S<sup>3\*</sup>; reduced/ox, ag.

2Cs + Br<sub>2</sub> - 2CsBr Cs<sup>6</sup> to Cs<sup>1+</sup>: oxidized/red. sg. Br<sup>0</sup> to Br<sup>1+</sup>; reduced/ox. ag.

3Mg + N<sub>2</sub> → Mg<sub>3</sub>N<sub>2</sub> Mg<sup>0</sup> to Mg<sup>2+</sup>; oxidized/red. ag. N<sup>0</sup> to N<sup>2+</sup>; reduced/ox. ag.

4Fe + 3O<sub>2</sub> → 2Fe<sub>2</sub>O<sub>3</sub> Fe<sup>6</sup> to Fe<sup>3\*</sup>; oxidized/red. ag. O<sup>6</sup> to O<sup>1\*</sup>; reduced/ox. ag.

Cl<sub>2</sub> + 2NaBr → 2NaCl + Br<sub>2</sub> Cl<sup>0</sup> to Cl<sup>1</sup>; reduced/ox. ag. Br<sup>1</sup> to Br<sup>0</sup>; oxidized/red. ag.

Si + 2F<sub>2</sub> → SiF<sub>4</sub> Si<sup>6</sup> to Si<sup>4</sup>; exidical red, ag F<sup>5</sup> to F<sup>1</sup>; reduced/ox, ag.

Mg + 2HCl → MgCl<sub>2</sub> + H<sub>2</sub> Mg<sup>0</sup> to Mg<sup>2+</sup>; oxidized/red. ag. H<sup>1+</sup> to H<sup>0</sup>; reduced/o.a.

10. 2Na + 2H<sub>2</sub>O → 2NaOH + H<sub>2</sub> Na<sup>0</sup> to Na<sup>1\*</sup>; oxidized/r.a. H<sup>1\*</sup> to H<sup>0</sup>; reduced/o.a.

Give the oxidation number of each kind of atom or ion.

a. sulfate b. Sn c. S<sup>2-</sup> d. Fe<sup>3+</sup> e. Sn<sup>4+</sup> f. nitrate g. ammonium 2- 0 2- 3+ 4+ 1- 1+

12. Calculate the oxidation number of chromium in each of the following.

a.  $Cr_2O_3$  b.  $Na_2Cr_2O_7$  c.  $CrSO_4$  d. chromate e. dichromate 3+ 6+ 2+ -74 6+

13. Use the changes in oxidation numbers to determine which elements are oxidized and which are reduced in these reactions. (Note: it is not necessary to use balanced equations)

a. C + H<sub>2</sub>SO<sub>4</sub> → CO<sub>2</sub> + SO<sub>2</sub> + H<sub>2</sub>O C<sup>0</sup> to C<sup>4+</sup>; oxidized S<sup>6+</sup> to S<sup>4+</sup>; reduced

b. HNO<sub>3</sub> + HI -- NO + I<sub>2</sub> + H<sub>2</sub>O N<sup>5+</sup> to N<sup>2+</sup>; reduced I<sup>1-</sup> to I<sup>0</sup>; exidized

c. KMnO<sub>4</sub> + HCl  $\longrightarrow$  MnCl<sub>2</sub> + Cl<sub>2</sub> + H<sub>2</sub>O + KCl Mn<sup>2</sup> to Mn<sup>2</sup>; reduced Cl to Cl<sup>2</sup>; exidized

d. Sb + HNO<sub>3</sub>  $\longrightarrow$  Sb<sub>2</sub>O<sub>3</sub> + NO + H<sub>2</sub>O Sb<sup>0</sup> to Sb<sup>3+</sup>; oxidized N<sup>5+</sup> to N<sup>2+</sup>; red.

14. For each reaction in problem 13, identify the oxidizing agent and reducing agent.

a. oxidizing agent: sulfur reducing agent: carbon

b. oxidizing agent: nitrogen reducing agent: iodine

oriefizing agent, manganese reducing agent: chloride.

d. oxidizing agent: nitrogen reducing agent: antimony

# **Redox Reduction Oxidation Reactions Answer**

J Ma

Redox Reduction Oxidation Reactions Answer:

### Redox Reduction Oxidation Reactions Answer Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has be more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Redox Reduction Oxidation Reactions Answer**," written by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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