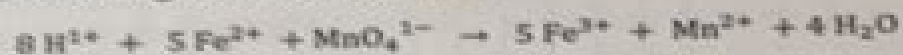


Pre-Lab Assignment: Redox Titration of Bleach

A student attempting to determine the mass percent of iron in a sample of ore reacts the ore with strong acid to convert the Fe to Fe^{2+} . Then the Fe^{2+} is titrated with MnO_4^{1-} according to the following balanced equation:



The student used 1.1081 grams of the ore and needed 39.32 mL of 0.03190 M KMnO_4 (the source of the MnO_4^{1-}) to complete the titration.

1. How many moles of permanganate ion, MnO_4^{1-} , did the student add to the ore?
2. What mass of iron was found to be present in the ore?
3. Determine the mass percent of the iron in the ore.

Redox Titration Pre Lab Questions Answers

M Tight



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