

## Chem 12 Electrochemistry : Quiz 2 - 120

### Part 1

1) A particular reaction has a predicted  $E^\circ$  of -2.25 V. This indicates that the reaction would :

- a) occur rapidly                      b) produce energy                      c) go to completion  
d) be non-spontaneous

2) Corrosion of a metal means that the metal :

- a) acts as a cathode                      b) undergoes oxidation  
c) undergoes reduction                      d) acts as an oxidizing agent

3) Zinc metal is prepared by electrolysis as follows:



What is the minimum voltage necessary for the electrolysis reaction to proceed ?

- a) 2.12                      b) 1.58                      c) 1.96                      d) 3.48

4) Which metal would provide the best cathodic protection for iron pipes ?

- a) Ag                      b) Sn                      c) Mg                      d) Cu

5) Which half-reaction best describes what is taking place at the anode during the electrolysis of a 1.0 M NaCl solution ?

- a)  $\text{Na}^{+} + \text{e}^{-} \rightarrow \text{Na}$                       b)  $2\text{Cl}^{-} \rightarrow \text{Cl}_2(\text{g}) + 2\text{e}^{-}$   
c)  $2\text{H}^{+}(10^{-7}\text{M}) + 2\text{e}^{-} \rightarrow \text{H}_2(\text{g})$                       d)  $\text{H}_2\text{O} \rightarrow 1/2 \text{O}_2(\text{g}) + 2\text{H}^{+}(10^{-7}\text{M}) + 2\text{e}^{-}$

6) Which one of the following converts chemical energy to electrical energy ?

- a) Industrial production of aluminum                      b) Nickel plating  
c) Discharging a "dry" cell                      d) Production of Na(s)

7) Which is the unit for the energy per coulomb of electrons ?

# Redox And Electrochemistry Quiz Answers

**Richard Bailey**



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## **Table of Contents Redox And Electrochemistry Quiz Answers**

1. Understanding the eBook Redox And Electrochemistry Quiz Answers
  - The Rise of Digital Reading Redox And Electrochemistry Quiz Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Redox And Electrochemistry Quiz Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Redox And Electrochemistry Quiz Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Redox And Electrochemistry Quiz Answers
  - Personalized Recommendations
  - Redox And Electrochemistry Quiz Answers User Reviews and Ratings
  - Redox And Electrochemistry Quiz Answers and Bestseller Lists
5. Accessing Redox And Electrochemistry Quiz Answers Free and Paid eBooks
  - Redox And Electrochemistry Quiz Answers Public Domain eBooks
  - Redox And Electrochemistry Quiz Answers eBook Subscription Services
  - Redox And Electrochemistry Quiz Answers Budget-Friendly Options
6. Navigating Redox And Electrochemistry Quiz Answers eBook Formats

- ePub, PDF, MOBI, and More
- Redox And Electrochemistry Quiz Answers Compatibility with Devices
- Redox And Electrochemistry Quiz Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Redox And Electrochemistry Quiz Answers
  - Highlighting and Note-Taking Redox And Electrochemistry Quiz Answers
  - Interactive Elements Redox And Electrochemistry Quiz Answers
- 8. Staying Engaged with Redox And Electrochemistry Quiz Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Redox And Electrochemistry Quiz Answers
- 9. Balancing eBooks and Physical Books Redox And Electrochemistry Quiz Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Redox And Electrochemistry Quiz Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Redox And Electrochemistry Quiz Answers
  - Setting Reading Goals Redox And Electrochemistry Quiz Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Redox And Electrochemistry Quiz Answers
  - Fact-Checking eBook Content of Redox And Electrochemistry Quiz Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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