

CHAPTER-8
REDOX REACTIONS

<u>oxidation</u>	<u>reduction</u>
1. Addition of oxygen	1. Removal of oxygen
2. Removal of hydrogen	2. Addition of hydrogen
3. Addition of an electronegative element	3. Removal of an electronegative element
4. Removal of an electropositive element	4. Addition of an electropositive element
5. Loss of electron	5. Gain of electron

Oxidation number denotes the oxidation state of an element in a compound ascertained according to a set of rules formulated on the basis that electron in a covalent bond belongs entirely to more electronegative element.

Calculation of oxidation number-

1. O. S. of all the elements in their elemental form (in standard state) is taken as zero. O. S. of elements in Cl_2 , F_2 , O_2 , P_4 , O_3 , Fe(s) , H_2 , N_2 , C(graphite) is zero.
2. Common O. S. of elements of group one (1^{st}) is one. Common O. S. of elements of group two (2^{nd}) is two.
3. For ions composed of only one atom, the oxidation number is equal to the charge on the ion.
4. The oxidation number of oxygen in most compounds is -2 . While in peroxides (e.g., H_2O_2 , Na_2O_2), each oxygen atom is assigned an oxidation number of -1 , in superoxides (e.g., KO_2 , RbO_2) each oxygen atom is assigned an oxidation number of $-(\frac{1}{2})$.
5. In oxygen difluoride (OF_2) and dioxygen difluoride (O_2F_2), the oxygen is assigned an oxidation number of $+2$ and $+1$, respectively.
6. The oxidation number of hydrogen is $+1$ but in metal hydride its oxidation no. is -1 .
7. In all its compounds, fluorine has an oxidation number of -1 .
8. The algebraic sum of the oxidation number of all the atoms in a compound must be zero.

Redox Reactions Study Guide Answer Key

Thomas Griffiths



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