

- Which subatomic particles are located in the nucleus of a neon atom?
 - electrons and positrons
 - electrons and neutrons
 - protons and neutrons
 - protons and electrons
- Which statement best describes the nucleus of an aluminum atom?
 - It has a charge of +13 and is surrounded by a total of 10 electrons.
 - It has a charge of +13 and is surrounded by a total of 13 electrons.
 - It has a charge of -13 and is surrounded by a total of 10 electrons.
 - It has a charge of -13 and is surrounded by a total of 13 electrons.
- Which statement concerning elements is true?
 - Different elements must have different numbers of isotopes.
 - Different elements must have different numbers of neutrons.
 - All atoms of a given element must have the same mass number.
 - All atoms of a given element must have the same atomic number.
- What are the characteristics of a neutron?
 - It has no charge and no mass.
 - It has no charge and a mass of 1 amu.
 - It has a charge of +1 and no mass.
 - It has a charge of +1 and a mass of 1 amu.
- Which sequence represents a correct order of historical developments leading to the modern model of the atom?
 - the atom is a hard sphere → most of the atom is empty space → electrons exist in orbitals outside the nucleus
 - the atom is a hard sphere → electrons exist in orbitals outside the nucleus → most of the atom is empty space
 - most of the atom is empty space → electrons exist in orbitals outside the nucleus → the atom is a hard sphere
 - most of the atom is empty space → the atom is a hard sphere → electrons exist in orbitals outside the nucleus
- Which conclusion is based on the "gold foil experiment" and the resulting model of the atom?
 - An atom is mainly empty space, and the nucleus has a positive charge.
 - An atom is mainly empty space, and the nucleus has a negative charge.
 - An atom has hardly any empty space, and the nucleus has a positive charge.
 - An atom has hardly any empty space, and the nucleus has a negative charge.
- Subatomic particles can usually pass undeflected through an atom because the volume of an atom is composed of
 - an uncharged nucleus
 - largely empty space
 - neutrons
 - protons
- What is the net charge on an ion that has 9 protons, 11 neutrons, and 10 electrons?
 - 1+
 - 2+
 - 1-
 - 2-
- What is the total charge of the nucleus of a carbon atom?
 - 6
 - 0
 - +6
 - +12
- Which notation represents an atom of sodium with an atomic number of 11 and a mass number of 24?
 - ${}_{11}^{24}\text{Na}$
 - ${}_{24}^{11}\text{Na}$
 - ${}_{11}^{24}\text{Na}$
 - ${}_{24}^{11}\text{Na}$
- Compared to an atom of phosphorus-31, an atom of sulfur-32 contains
 - one less neutron
 - one less proton
 - one more neutron
 - one more proton
- Which subatomic particle has a negative charge?
 - proton
 - electron
 - neutron
 - positron
- Compared to a proton, an electron has
 - a greater quantity of charge and the same sign
 - a greater quantity of charge and the opposite sign
 - the same quantity of charge and the same sign
 - the same quantity of charge and the opposite sign
- How many electrons are in an Fe^{2+} ion.
 - 24
 - 26
 - 28
 - 56
- Which two particles make up most of the mass of a hydrogen-2 atom?
 - electron and neutron
 - electron and proton
 - proton and neutron
 - proton and positron
- Which two particles each have a mass approximately equal to one atomic mass unit?
 - electron and neutron
 - electron and positron
 - proton and electron
 - proton and neutron
- What is the mass number of an atom that has six protons, six electrons, and eight neutrons?
 - 6
 - 12
 - 14
 - 20
- What is the total number of neutrons in the nucleus of a neutral atom that has 19 electrons and a mass number of 39?
 - 19
 - 20
 - 39
 - 58
- In which list are the elements arranged in order of increasing atomic mass?
 - Cl, K, Ar
 - Fe, Co, Ni
 - Te, I, Xe
 - Ne, F, Na

Regents Review Atomic Structure Answers

Richard Bailey



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