

III. Determining the pH of various salt solutions.

During the experiment you would obtain a small amount of aqueous salts, and would measure their pH. You can predict whether a salt solution will be acidic, basic, or neutral following four situations:

1. A strong acid and a strong base will react to form a **neutral salt solution**.
2. A strong acid and a weak base yield a weakly **acidic salt solution**.
3. A weak acid and a strong base yield a weakly **basic salt solution**.
4. A weak acid plus a weak base can yield either **an acidic or basic solution**. This is the most complex of the four types of reactions. To predict whether a particular combination will be acidic or basic, tabulated K values of the conjugates must be compared. 4

Table 3. Determining the pH of various salt solutions

Salt	Name of salt	Salt type (acidic/basic)	Measured pH	Explanation for the difference between measured pH and 7.0
NaCl	Sodium chloride	Neutral salt	7.23	
<u>NaC₂H₃O₂</u>	Sodium acetate	Basic salt	9.41	
<u>NaHCO₃</u>	Sodium bicarbonate	Basic salt	9.29	
<u>Na₂CO₃</u>	Sodium carbonate	Basic salt	10.67	
<u>NH₄Cl</u>	Ammonium chloride	Acidic salt	5.79	
NH ₄ C ₂ H ₃ O ₂	Ammonium acetate	Basic salt	6.95	

How to write an explanation for the difference between measured pH and 7.0? You need to provide an equation for hydrolysis of each salt. This will confirm that there is a formation of OH⁻ or H₃O⁺

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1. Understanding the eBook Ph Of Salt Solutions Worksheet Answers If8767
 - The Rise of Digital Reading Ph Of Salt Solutions Worksheet Answers If8767
 - Advantages of eBooks Over Traditional Books
2. Identifying Ph Of Salt Solutions Worksheet Answers If8767
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ph Of Salt Solutions Worksheet Answers If8767
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ph Of Salt Solutions Worksheet Answers If8767
 - Personalized Recommendations
 - Ph Of Salt Solutions Worksheet Answers If8767 User Reviews and Ratings
 - Ph Of Salt Solutions Worksheet Answers If8767 and Bestseller Lists
5. Accessing Ph Of Salt Solutions Worksheet Answers If8767 Free and Paid eBooks
 - Ph Of Salt Solutions Worksheet Answers If8767 Public Domain eBooks
 - Ph Of Salt Solutions Worksheet Answers If8767 eBook Subscription Services
 - Ph Of Salt Solutions Worksheet Answers If8767 Budget-Friendly Options
6. Navigating Ph Of Salt Solutions Worksheet Answers If8767 eBook Formats
 - ePub, PDF, MOBI, and More
 - Ph Of Salt Solutions Worksheet Answers If8767 Compatibility with Devices
 - Ph Of Salt Solutions Worksheet Answers If8767 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ph Of Salt Solutions Worksheet Answers If8767
 - Highlighting and Note-Taking Ph Of Salt Solutions Worksheet Answers If8767
 - Interactive Elements Ph Of Salt Solutions Worksheet Answers If8767
8. Staying Engaged with Ph Of Salt Solutions Worksheet Answers If8767

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ph Of Salt Solutions Worksheet Answers If8767
9. Balancing eBooks and Physical Books Ph Of Salt Solutions Worksheet Answers If8767
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ph Of Salt Solutions Worksheet Answers If8767
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ph Of Salt Solutions Worksheet Answers If8767
- Setting Reading Goals Ph Of Salt Solutions Worksheet Answers If8767
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ph Of Salt Solutions Worksheet Answers If8767
- Fact-Checking eBook Content of Ph Of Salt Solutions Worksheet Answers If8767
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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