

## pH Scale

You will not be given any of these formulae!!



$$\text{pH} = -\log [\text{H}^+]$$

$$\text{pOH} = -\log [\text{OH}^-]$$

$$\text{pH} + \text{pOH} = 14$$

$$K_w = [\text{H}^+][\text{OH}^-] = 1 \times 10^{-14}$$

### Problems

- 1) What is the pH of a solution with an  $\text{H}^+$  ion concentration of  $1 \times 10^{-6} \text{ M}$ ?
- 2) What is the pOH of a solution with an  $\text{OH}^-$  ion concentration of  $1 \times 10^{-3} \text{ mol/L}$ ?
- 3) What is the pOH of a solution with a pH of 9?
- 4) What is the pH of a solution with a pOH of 3?
- 5) What happens to the concentration of hydronium ion as you go from pH 5 to pH 2?  
M and E!

# Ph And Poh Continued Problems With Answers

**Clyde Metz**



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