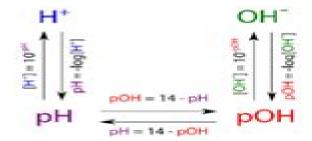
You will not be given any of these formulae!!



pH = -log [H'] pOH = -log [OH] pH + pOH = 14 $Kw = [H'] [OH] = 1 \times 10^{14}$

Problems

- What is the pH of a solution with an H° ion concentration of 1 x 10° M?
- What is the pOH of a solution with an OH ion concentration of 1 x 10³ mol/L?
- 3) What is the pOH of a solution with a pH of 9?
- 4) What is the pH of a solution with a pOH of 3?
- 5) What happens to the concentration of hydronium ion as you go from pH 5 to pH 2?
 M and E!

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