

Acids & Bases: Conj. Acid-Base Pairs & "pH block" Calculations Skill Review

Skill - The student will be able to identify the acid, base, conjugate-acid, and conjugate-base in an acid - base reaction.

Label the acid and base on the reactant side of each of the following equations and label the products as the conjugate-acid and conjugate-base.



Fill in the following table

	Acid	Base	Conjugate Acid	Conjugate Base	Equation
3.	HNO_3	H_2O			$\text{HNO}_3 + \text{H}_2\text{O} \rightleftharpoons \text{NO}_3^- + \text{H}_3\text{O}^+$
4.			H_2O	CO_3^{2-}	

5. Write an equation that shows the reaction of ammonia (NH_3) with hydrobromic acid (HBr). Label the acid, base, conjugate-acid & conjugate base.

Skill - Student will be able to calculate with the $[\text{H}^+]$ or $[\text{OH}^-]$ from the ion product (K_w) or pH from pOH (one-step pH block calculation)

6. Calculate the $[\text{OH}^-]$ of a solution with a $[\text{H}^+]$ of 1.00×10^{-10} .

7. Calculate the $[\text{H}^+]$ calculation of a solution with a pOH of 8.50.

8. Calculate the pOH of a solution with a pH of 8.63.

9. Calculate the pOH of a solution with a $[\text{OH}^-]$ of 5.62×10^{-14} .

Ph Calculations Answer Key

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