

Purpose: To study the ideal gas law and apply the first law of thermodynamics to analyze behavior of gas. Go to <http://phet.colorado.edu/en/simulation/gas-properties> and run the simulation

Procedure**PART A: THE NATURE OF THE IDEAL GAS LAW**

1: What happens to the pressure in the chamber if you heat the gas?

2: What happens to the temperature in the chamber if you compress the gas?

3: Locate the "Constant Parameter" section of the on-screen control panel. Lock the temperature. Use "Scubie" to slowly compress the gas. What happens to temperature, and what action is taken (by the simulation) to maintain constant temperature?

4: Lock the pressure. (Doing so releases the lock on temperature.) Add heat. What happens to pressure, and what action is taken (by the simulation) to maintain constant pressure?

PART B: Kinetic Molecular Theory and the First Law of Thermodynamics

For each of the following four steps record the average pressure (in atm) and the volume

Table of Pressure and Volume for each step

Step	1	2	3	4
Pressure (atm)				
Volume (m^3)				

1. Add 300 molecules of the heavy species to the chamber and let the temperature and pressure come to equilibrium.
 - a. Using Avogadro's number express the amount of gas in terms of moles.
 - b. Calculate the internal energy (U) of the heavy oxygen gas in joules .
 - c. The container is rectangular in shape with an area of $2.81 \times 10^{-17} \text{ m}^2$. Using the ruler tool (in measurement tools) measure the interior length of the container and report measurement in meters.
Length: _____ m
 - d. Calculate the volume of the container (in m^3)
 - e. Create different volumes and record the average pressure (in atm) and the volume in the table above for step 1.

Phet Exploring Gas Laws Lab Answers

Sir James Augustus Henry Murray



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