

## Stoichiometry Review

1. The combustion of propane occurs via the reaction:



How many grams of oxygen are required to burn completely 10.0 g of propane?

$$\left( \frac{10.0 \text{ g C}_3\text{H}_8}{44.09652 \text{ g/mol}} \right) \left( \frac{5 \text{ mol O}_2}{1 \text{ mol C}_3\text{H}_8} \right) \left( 31.9988 \text{ g/mol} \right) = 36.3 \text{ g O}_2$$

2. Small quantities of chlorine can be prepared in the laboratory by the reaction:



What volume of chlorine at STP can be prepared from 100.0 g of manganese dioxide?

$$\left( \frac{100.0 \text{ g MnO}_2}{86.9368 \text{ g/mol}} \right) \left( \frac{1 \text{ mol Cl}_2}{1 \text{ mol MnO}_2} \right) \left( 22.4 \text{ L/mol} \right) = 25.8 \text{ L Cl}_2$$

3. The most common ore of arsenic is mispickel, FeSAs. Upon heating this ore, free arsenic is obtained. How many grams of FeSAs are required to produce 10.0 g of arsenic according to the following reaction?



$$\left( \frac{10.0 \text{ g As}}{74.9216 \text{ g/mol}} \right) \left( \frac{1 \text{ mol FeSAs}}{1 \text{ mol As}} \right) \left( 162.8346 \text{ g/mol} \right) = 21.7 \text{ g FeSAs}$$

# Pg 31section Review Stoichiometry Answers

**R Barnett**



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