

Calculating pH of Acids – Practice

Instructions: Use the equations for pH to complete the problems below.

- 1) What is the pH of a 0.0266 M hydrobromic acid (HBr) solution?

$$\text{pH} = -\log[\text{H}^+] \rightarrow \text{pH} = -\log[0.0266] \rightarrow \boxed{\text{pH} = 1.58}$$

- 2) Calculate the hydrogen ion concentration of a solution with a pH of 6.3.

$$[\text{H}^+] = 10^{-\text{pH}} \rightarrow [\text{H}^+] = 10^{-6.3} \rightarrow \boxed{[\text{H}^+] = 5.01 \times 10^{-7}}$$

- 3) Calculate the pH of a 0.004 M hydrofluoric acid (HF) solution?

$$\text{pH} = -\log[\text{H}^+] \rightarrow \text{pH} = -\log[0.004] \rightarrow \boxed{\text{pH} = 2.40}$$

- 4) What is the hydrogen ion concentration of an acid with a pH of 2.21?

$$[\text{H}^+] = 10^{-\text{pH}} \rightarrow [\text{H}^+] = 10^{-2.21} \rightarrow \boxed{[\text{H}^+] = 6.17 \times 10^{-3}}$$

- 5) Calculate the pH of a 1.14 L solution that contains 0.19 moles of HCl.

$$[\text{H}^+] = \frac{\text{moles of solute}}{\text{liters of solvent}} \rightarrow \frac{0.19 \text{ mol}}{1.14 \text{ L}} \rightarrow 0.167 \text{ M}$$

$$\text{pH} = -\log[\text{H}^+] \rightarrow \text{pH} = -\log[0.167] \rightarrow \boxed{\text{pH} = 0.78}$$

Ph Review Problems Answers Chemfiesta

RD Boyd



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web in a large microwave safe bowl microwave the cream cheese for 10 seconds add the sugar and vanilla and combine until the mixture is smooth and creamy slowly mix in the heavy cream and milk until fully combined and sugar is dissolved pour base into an empty creami pint

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web in a small mixing bowl or large liquid measuring cup combine instant pudding mix and milk whisk together until pudding mix dissolves add heavy cream and vanilla extract whisking just to combine transfer mixture to a ninja creami pint being careful not to exceed the max fill line and close the lid

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web jun 8 2023 this ninja creami vanilla ice cream recipe is made with cream cheese sugar vanilla whole milk and heavy cream it s the perfect base recipe it customize