

Chemical 101

pH and pOH

Name: _____
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Introduction Questions

1. How are we going to get ready to learn about pH. What are general questions asking the suggested type of a solution. Information we're interested in: say pH 2.5.
a. Take the log with that parenthesis. How about $\log_{10} 2.5$
b. How does the suggested value answer you. Not the $\log_{10} 2.5$ value
c. Find the power of 10 that $\log_{10} 2.5 = 0.40$
d. Calculate the following:
i. $\log_{10} 1000$ ii. $\log_{10} 100$ iii. $\log_{10} 10$
 3.0 *2.0* *1.0*

Information: Calculating pH

Many properties based on pH^+ , and many people know that acidity is low, basicity is high, but that very many calculated what acidity is measured. For pH has standard definition: $\text{pH} = -\log_{10} [\text{H}^+]$ for a solution.

Simple pH is solution can range from 0 to 14. It's equivalent for calculating pH is:

$$\text{pH} = -\log_{10} [\text{H}^+]$$

The suggested pH is a solution for the concentration of $[\text{H}^+]$ in molarity of $10^{-\text{pH}}$. A pH of 7 is neutral solution. Below 7 is acidic and above 7 is basic.

Critical Thinking Questions

1. An aqueous solution has a concentration of hydrogen ions of 1.0×10^{-4} M. What is the pH?
 $\text{pH} = -\log_{10} 1.0 \times 10^{-4} = 4.0$
2. An aqueous solution has a concentration of 10^{-10} M of H^+ .
 a. What is the concentration of H^+ ions?
 10^{-10} M
 b. Calculate the pH of this solution.
 $\text{pH} = -\log_{10} 10^{-10} = 10$

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