CRITICAL POINT - The temperature above which a substance will always be a gas regardless of the pressure.

Pressure (atm)

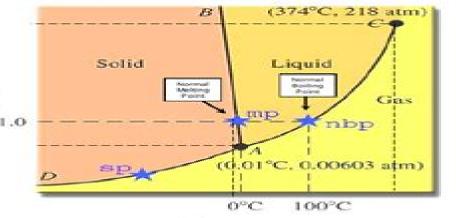
- NOTE:
  - The solid phase is more dense than the liquid phase.
  - The line between the solid and gas phases is the equilibrium of solid and gas phases at that specific pressure and temperature, i.e. a curve of all the deposition/sublimation points.
  - The line between the solid and liquid phases is the equilibrium of solid and liquid phases at that specific pressure and temperature, i.e. a curve of all the freezing/melting points.
  - The line between the liquid and gas phases is the equilibrium of liquid and gas phases at that specific pressure and temperature, i.e. a curve of all the vaporization/condensation points.
- Melting Point (Freezing Point) The temperature at which the solid and liquid phases of a substance are in equilibrium at atmospheric pressure.
  - Normal Melting Point (Freezing Point) The temperature at which the solid changes to a liquid at Standard Pressure (1.00 atm = 760 mmHg = 760 torr = 101.325 kPa)
- Boiling Point (Condensation Point) The temperature at which the vapor pressure of a liquid is equal
  to the pressure on the liquid.
  - Normal Boiling Point (Condensation Point) The temperature at which the vapor pressure of a liquid is equal to Standard Pressure (1.00 atm = 760 mmHg = 760 torr = 101.325 kPa)

## Phase Diagram for Water

For water, the liquid phase is more dense than the solid phase due to hydrogen bonding.

liquid, and gas phases exist

simultaneously.



Temperature

# **Phase Change Diagram Ws Answer**

Daniel J. Findley, Christopher M. Cunningham, Thomas H. Brown Jr, Lorraine M. Cahill, Guangchuan Yang, Leta F. Huntsinger

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