#### NEUTRALIZATION

The reaction of an acid with a base is called a "neutralization" reaction. There are two types that are emphasized here.

TYPE I.	Acid	+	Basic Hydroxide	0.00	H <sub>2</sub> O	:=	Salt (ionic compound)
	EXAMPLES (balance these equations:)						
	HCI	+	NaOH	12	$H_2O$	+	NaCl
	HCl	+	Ca(OH) <sub>2</sub>	100	$H_2O$	*	CaCl <sub>2</sub>
	H <sub>2</sub> SO <sub>4</sub>	+	кон	177	H <sub>2</sub> O	+	$K_2SO_4$
	H <sub>2</sub> SO <sub>4</sub>	+	Ba(OH) <sub>2</sub>	: <del>**</del>	$H_2O$	+	BaSO <sub>a</sub>
	$H_3PO_4$	+	NaOH	1000	$H_2O$	+	$Na_3PO_4$
	$H_3PO_4$	+ -	Ca(OH) <sub>2</sub>	-	$H_2O$	-	$Ca_{3q}PO_4)_2$
TYPE 2.	Acid	+	Ammonia	0.400	Amm	oniu	nm Salt (ionic compound)
	EXAMPLES (balance these equations:)						
	HCI	+	$NH_3$	277	NH <sub>4</sub> C	i	
	H <sub>2</sub> SO <sub>4</sub>	+	NH <sub>3</sub>	100	(NH <sub>4</sub> )	so.	
	H,PO4	+	NH <sub>3</sub>	875	(NH <sub>4</sub> )	,PO	

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**Amar Nath Bhutani** 

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