



Percentage Yield

Sample Problem H Solution, *continued*



Theoretical yield

$$\begin{aligned} 36.8 \text{ g C}_6\text{H}_6 &\times \frac{1 \text{ mol C}_6\text{H}_6}{78.12 \text{ g C}_6\text{H}_6} \times \frac{1 \text{ mol C}_6\text{H}_5\text{Cl}}{1 \text{ mol C}_6\text{H}_6} \times \frac{112.56 \text{ g C}_6\text{H}_5\text{Cl}}{1 \text{ mol C}_6\text{H}_5\text{Cl}} \\ &= 53.0 \text{ g C}_6\text{H}_5\text{Cl} \end{aligned}$$

Percentage yield

$$\text{percentage yield C}_6\text{H}_5\text{Cl} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

$$\text{percentage yield} = \frac{38.8 \text{ g}}{53.0 \text{ g}} \times 100 = 73.2\%$$



Section 3 Limiting Reactants And Percentage Yield

**Theodore L. Brown, H. Eugene LeMay
Jr., Bruce E. Bursten, Catherine
Murphy, Patrick Woodward, Steven
Langford, Dalius Sagatys, Adrian
George**

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