

**52****ELECTRON ARRANGEMENT IN ATOMS****Section Review****Objectives**

- Describe how to write the electron configuration for an atom
- Explain why the actual electron configurations for some elements differ from those predicted by the Aufbau principle

**Vocabulary**

- electron configurations
- Aufbau principle
- Pauli exclusion principle
- Hund's rule

**Part A Completion**

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

The ways in which electrons are arranged around the nuclei of atoms are called 1. The 2 describes the sequence in which orbitals are filled. The various orbitals within a sublevel of a principle energy level are always of 3 energy. The 4 principle states that a maximum of only 5 electrons can occupy each orbital. To occupy the same orbital, two electrons must have 6 spins. Hund's rule states that the electrons pair up only after each orbital in a sublevel is occupied by 7. When using the shorthand method for showing the electron configuration of an atom, 8 are used to indicate the number of 9 occupying each sublevel.

Correct electron configurations can be obtained by using the Aufbau diagram for the elements up to and including vanadium. 10 and copper are exceptions to the Aufbau principle.

1. \_\_\_\_\_
2. \_\_\_\_\_
3. \_\_\_\_\_
4. \_\_\_\_\_
5. \_\_\_\_\_
6. \_\_\_\_\_
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9. \_\_\_\_\_
10. \_\_\_\_\_

# Section 5 2 Electron Arrangement In Atoms Answer Key

**Fujia Yang, Joseph H Hamilton**



## Section 5 2 Electron Arrangement In Atoms Answer Key:

*Understanding General Chemistry* Atef Korchef, 2022-03-07 *Understanding General Chemistry* details the fundamentals of general chemistry through a wide range of topics relating the structure of atoms and molecules to the properties of matter. Written in an easy to understand format with helpful pedagogy to fuel learning the book features main objectives at the beginning of each chapter, get smart sections and check your reading section at the end of each chapter. The text is filled with examples and practices that illustrate the concepts at hand. In addition, a summary and extensive MCQs exercises and problems with the corresponding answers and explanations are readily available. Additional features include Alerts students to common mistakes and explains in simple ways and clear applications how to avoid these mistakes. Offers answers and comments alongside sample problems enabling students to self evaluate their skill level. Includes powerful methods, easy steps, simple and accurate interpretations and engaging applications to help students understand complex principles. Provides a bridge to more complex topics such as solid state chemistry, organometallic chemistry, chemistry of main group elements, inorganic chemistry and physical chemistry. This introductory textbook is ideal for chemistry courses for non science majors as well as health sciences and preparatory engineering students.

**General Chemistry** Raymond E. Davis, Kenneth W. Whitten, 1996 Are you looking for the key to success in your chemistry class? In *CHEMISTRY* you will find a strong molecular reasoning focus, problem solving exercises and an innovative online homework management system that will prepare you for any challenge you might encounter. The textbook is filled with learning aids that will help you master concepts of the course.

**Fundamentals of Chemistry** Ralph A. Burns, 1995      [Spectra of Atoms and Molecules](#) Peter F. Bernath, 2005-04-21

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