

Chapter 8 Solutions, Acids, and Bases

Section 8.2 Solubility and Concentration**(pages 235–239)**

This section explains solubility, the factors affecting solubility, and different ways of expressing the concentration of a solution.

Reading Strategy (page 235)

Previewing: Before you read the section, rewrite the topic headings as *how, why,* and *what* questions. As you read, write an answer to each question. For more information on this reading strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

Question	Answer
What is solubility?	
	Solvent, temperature, and pressure

Solubility (pages 235–239)

- Define solubility. _____
- List the following solutes in order from most soluble to least soluble in water: table salt, baking soda, table sugar.
 - _____
 - _____
 - _____
- Circle the letters that identify how solutions can be classified based on solubility.
 - unsaturated
 - desaturated
 - saturated
 - supersaturated
- Describe a saturated solution. _____
- A solution that has less than the maximum amount of solute that can be dissolved is called a(n) _____.
 - unsaturated
 - saturated
 - supersaturated
 - desaturated
- Is the following sentence true or false? It is impossible for a solution to contain more solute than the solvent can hold at a given temperature. _____

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