

11.1

DESCRIBING CHEMICAL REACTIONS

Section Review

Objectives

- 1. Represent the chemical equation
- 2. Explain how to interpret the equation
- 3. List the steps for writing a chemical equation

Vocabulary

- 1. chemical equation
- 2. chemical equation
- 3. reactant
- 4. coefficient
- 5. balanced equation

Part A Completion

Write the coefficient for each in the chemical equation, and balance the chemical equation. Write the name of each reactant and product.

_____ a chemical equation is always represented by a balanced _____
 The chemical equation is a chemical change called _____ that uses
 chemical substances (reactants) to make _____ products and energy.
 The law of conservation of _____ a chemical reaction neither creates nor
 destroys matter. _____ is given _____ as a chemical equation
 indicates that the high content of heat in the _____ is not sufficient to
 the system the system has energy added by burning the _____ of
 the system (the reactants).

Chemical equations are written with physical states (solid, liquid, gas, or aqueous) as a
 subscript. The equation is _____ if it is a solid _____ if it is a gas
 _____ if it is a liquid _____ if it is a solution. _____
 must be placed above a chemical equation to indicate that the
 reaction is reversible.

- 1. equation
- 2. reaction
- 3. products
- 4. energy
- 5. conservation
- 6. matter
- 7. heat energy
- 8. _____
- 9. _____
- 10. _____
- 11. _____
- 12. _____
- 13. _____
- 14. _____
- 15. _____

Section Review 1describing Chemical Reactions Answer Key

Ralph Thomas Myers



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Study Questions and Problems in Inorganic Chemistry Alexander Silverman, 1924

Solvent Effects and Chemical Reactivity Orlando Tapia, Juan Bertrán, 2006-04-11 This book gathers original contributions from a selected group of distinguished researchers that are actively working in the theory and practical applications of solvent effects and chemical reactions The importance of getting a good understanding of surrounding media effects on chemical reacting system is difficult to overestimate Applications go from condensed phase chemistry biochemical reactions in vitro to biological systems in vivo Catalysis is a phenomenon produced by a particular system interacting with the reacting subsystem The result may be an increment of the chemical rate or sometimes a decreased one At the bottom catalytic

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