REVIEW: Chapter 8 - Covalent Bonding

1. Typically, how many bonds & lone pairs will each have when covalently bonding?

a. Boron - 3 bonds: 0 lone pairs

d. lodine - 1 bond; 3 lone pairs

b. Silicone - 4 bonds; O lone pairs

e. Nitrogen - 3 bonds; 1 lone pair

c. Neon - 0 bonds; 4 lone pairs

f. Oxygen - 2 bonds; 2 lone pairs

2. Which elements do not follow the octet rule? (Hint - there are four we went over)

a. __H__ & __B__ never follow the octet rule.

b. __P__ & __5_ sometimes don't follow the octet rule.

3. What type of bond will each combination of atoms form?

a. F & N = 2 nonmetals, therefore covalent

d. B & CI - 2 nonmetals, therefore covalent

b. Al & O - 1 metal/1 nonm., therefore ionic

e. Ca & I - 1 metal/1 nonm., therefore ionic

c. Na & 5 - 1 metal/1 nonm., therefore ionic

f. Fe & Cu - 2 metals, therefore metallic

- 4. What is the difference between:
 - a. Molecules & Compounds?

Compounds can be made using all types of bonds (ionic, covalent or metallic), but molecules can only be made using covalent bonds

b. Monatomic elements & Diatomic molecules ?

Monatomic elements exist as single atoms; diatomic molecules are made up of two of the same atom

c. Structural formulas, Empirical formulas, & Molecular formulas?

All three are for covalent bonds only! Molecular show the actual number of atoms in the molecule; Empirical give the lowest whole number ratio; Sturctural formulas are not really formulas, they are drawing of the molecule.

d. Polar bonds & Non-polar bonds?

Both describe covalent bonds! In a polar bond the electrons are NOT shared equally; in a non-polar bond electrons are shared equally

5. Classify each of the following as an atom or a molecule:

a. Be, atom

b. CO, molec.

c. Na molec.

d. H₂O, molec.

e. Ne. atom

Section 8 Covalent Bonding Workbook Answers

VM Jensen

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