

Activity B (continued from previous page)

4. **Calculate:** Select the **Worksheet** tab. This tab helps you calculate the analyte concentration.
- Fill in the first set of boxes ("moles H_2SO_4 " and "moles NaOH ") based on the coefficients in the balanced equation. (If there is no coefficient, the value is 1.)
 - Record the appropriate volumes in the "mL NaOH " and "mL H_2SO_4 " boxes.
 - Record the concentration of the titrant in the M NaOH box.

Click **Calculate**. What is the concentration listed? _____

Click **Check**. Is this the correct concentration? _____

If you get an error message, revise your work until you get a correct value. (You may have to redo the titration if you do not have the correct volume of titrant.)

5. **Practice:** Perform the following titrations and determine the concentrations of the following solutions. In each experiment, list the volume of titrant needed to neutralize the analyte and the indicator used. Use the **Worksheet** tab of the Gizmo to calculate each analyte concentration. Include all units.

Titrant	Analyte	Indicator	Titrant volume	Analyte concentration
0.70 M KOH	HBr			
0.50 M HCl	$\text{Ca}(\text{OH})_2$			
0.80 M H_2SO_4	NaOH			

6. **Apply:** Once you know the concentration of a strong acid or a strong base, you can estimate its pH. Use $\text{pH} = -\log_{10}[\text{H}^+]$ to calculate the pH of each of the strong acid mystery solutions (**Mystery HBr** and **Mystery H_2SO_4**) based on the concentrations you determined in questions 4 and 5. Check your answers with the Gizmo. (Because dissociation is not always complete, your answers may vary slightly from values in the Gizmo.)

$[\text{H}_2\text{SO}_4] =$ _____ $\text{pH } \text{H}_2\text{SO}_4 =$ _____ $[\text{HBr}] =$ _____ $\text{pH HBr} =$ _____

7. **Apply:** For a strong base, the concentration of hydroxide ions $[\text{OH}^-]$ is roughly estimated to be the same as the concentration of the base. The pH of a strong base is found with the equation $\text{pH} = 14 + \log_{10}[\text{OH}^-]$. Based on their concentrations, find the pH of each of the strong bases. Check your answers with the Gizmo.

$[\text{Ca}(\text{OH})_2] =$ _____ $\text{pH } \text{Ca}(\text{OH})_2 =$ _____ $[\text{NaOH}] =$ _____ $\text{pH NaOH} =$ _____

Self Exploration Stoichiometry

Didier Musso



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