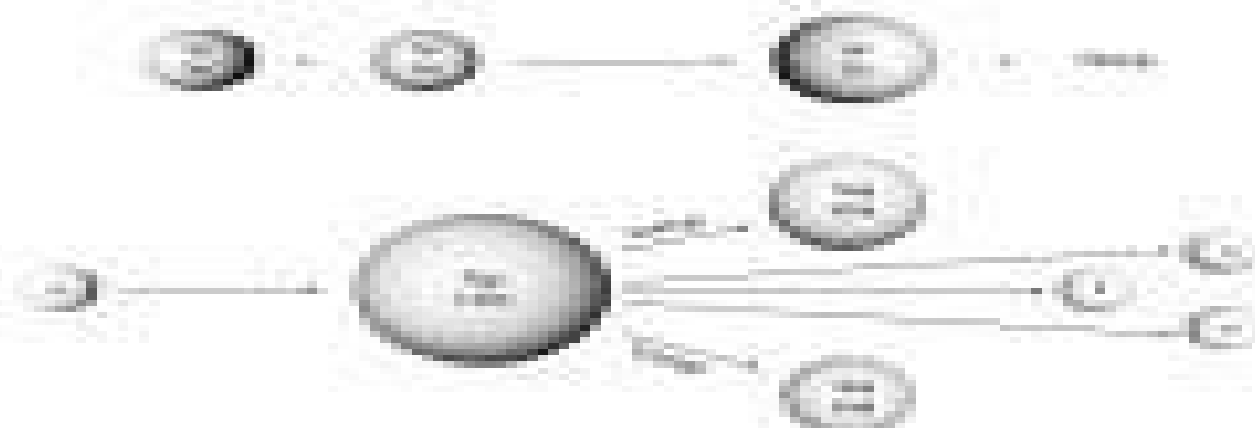




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## Nuclear Decay and Reactions

[illegible]

- a. The diagram shows two types of chemical reactions: combustion reactions and synthesis reactions. Label the types of chemical reactions for each diagram as given.
- b. Label reaction 1: Combustion Reaction      Reaction 2: Synthesis Reaction
- c. Write the balanced chemical equation for reaction 1's combustion. The reaction must be balanced and include physical states.
- d.  $2H_2 + O_2 \rightarrow 2H_2O$       e.  $2H_2 + O_2 \rightarrow 2H_2O$
- e. Write the balanced chemical equation for reaction 2's synthesis by applying the reaction stoichiometric ratios from the diagram.
- f. a.  $2H_2 + O_2 \rightarrow 2H_2O$       b.  $2H_2 + O_2 \rightarrow 2H_2O$
- g. a.  $2H_2 + O_2 \rightarrow 2H_2O$       b.  $2H_2 + O_2 \rightarrow 2H_2O$
- h.  $2H_2 + O_2 \rightarrow 2H_2O$       i.  $2H_2 + O_2 \rightarrow 2H_2O$
- i. Write the balanced chemical equation for reaction 2's synthesis by applying the reaction stoichiometric ratios from the diagram.
- j. Label each diagram with the reaction type: Combustion Reaction      Synthesis Reaction
- k. Label each diagram with the reaction type: Combustion Reaction      Synthesis Reaction

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12. Explain each species listed on the last classification table with an identifying signal and the nature of its signal.

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## Answers Key

**O García**



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