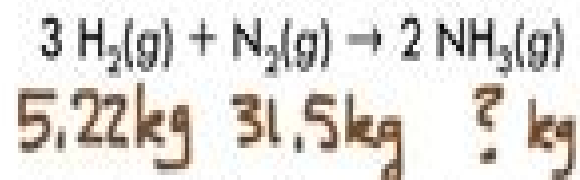


Limiting Reactant & Theoretical Yield

21

Ammonia can be synthesized by the reaction:



What is the theoretical yield of ammonia, in kg, that we can synthesize from 5.22 kg of H_2 and 31.5 kg of N_2 ?

$$5.22 \text{ kg H}_2 \times \frac{1 \text{ mol H}_2}{2.016 \text{ g H}_2} \times \frac{\text{mol NH}_3}{\text{mol H}_2} \times \frac{\text{g NH}_3}{\text{mol NH}_3}$$

Section 9 3 Limiting Reagent And Percent Yield Key

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