

## 9.1

## THE ARITHMETIC OF EQUATIONS

## SECTION REVIEW

**Objectives**

- Calculate the amount of reactants required or product formed in a chemical process.
- Interpret balanced chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

**Key Terms**

- stoichiometry

**Part A Completion**

Copy this completion exercise for class. Your understanding of the concepts and terms is clear when indicated in this exercise. Blank blanks are the completed work of others; absent phrases are none.

- The coefficients in a balanced chemical equation indicate \_\_\_\_\_ 1. \_\_\_\_\_ the relative number of \_\_\_\_\_ 2. \_\_\_\_\_ of reactants and products. All stoichiometric calculations begin with a \_\_\_\_\_ 3. \_\_\_\_\_ (molar \_\_\_\_\_ 4. \_\_\_\_\_) and \_\_\_\_\_ 5. \_\_\_\_\_ are converted in every reaction; moles, volumes, and representative particles may not be.
- In solving stoichiometric problems, conversion factors relating moles of reactants to \_\_\_\_\_ 6. \_\_\_\_\_ of products are used.
- If you assume \_\_\_\_\_ 7. \_\_\_\_\_, the equation also tells you about the volume of gas.

**Part B True-False**

(Quickly mark off the statements as to be always true, AC; sometimes true, ST; or never true, NT.)

- \_\_\_\_\_ 7. The coefficients in a balanced chemical equation can be used to derive molar ratios relating reactants to products.
- \_\_\_\_\_ 8. The coefficients in a balanced chemical equation tell the relative volumes of reactants and products, expressed in any suitable units of volume.
- \_\_\_\_\_ 9. To calculate the mass of a molecule in grams, you can use the molar mass and Avogadro's number.

# The Arithmetic Of Equations 9 1 Answer Key

**BM King**



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**Educational Times and Journal of the College of Preceptors**, 1908 *The IIT Foundation Series - Mathematics Class 10, 2/e*,

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**Progressive Arithmetic** William James Milne, 1906

**The Encyclopaedia of Pure Mathematics**, 1847

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