

## CHEMICAL EQUATIONS

A **chemical equation** is a simple and easy way to describe a chemical reaction. **Chemical equations use chemical formulas and other symbols instead of words to summarize a reaction.** A chemical equation tells you the substances you start with in a reaction and the substances that are formed in the end. The chemical equation for the formation of water is shown below.



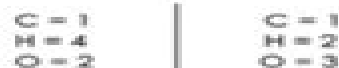
All chemical equations share a common structure. The substances you have at the beginning are called the **reactants**. The new substances formed at the end of the reaction are called the **products**. The formulas for the reactants are written on the left side of the equation, followed by an **arrow** (  $\rightarrow$  ) which means "turns into" or "yields". The formulas for the products are written on the right side of the equation. When there are two or more reactants or products, they are separated by plus signs.

The **law of conservation of mass** was first demonstrated by a scientist named Antoine Lavoisier in the late 1700s. The law of conservation of mass states that in a chemical reaction, **the total mass of the reactants must equal the total mass of the products**. In an **open system**, matter can enter or escape from the surroundings. For example, a match burning in the open air is an open system. You cannot measure the mass of all the reactants and products in an open system. In a **closed system**, matter cannot enter or escape into the surroundings. A match burning inside a sealed glass chamber is an example of a closed system. A closed system allows you to measure the mass of all the reactants and products in a reaction.

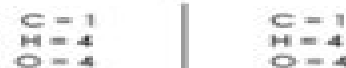
Because of the law of conservation of mass, a chemical equation must be **balanced**. An equation is balanced when the total number of atoms of the reactants equals the total number of atoms in the products. To balance a chemical reaction, you may have to use coefficients. A **coefficient** is a number placed in front of a chemical formula in an equation. The coefficient tells you how many molecules of a reactant or product take part in a reaction.



**unbalanced** equation



**balanced** equation



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