

# Chemistry Paper 1 revision

## Making salts

All salts contain a positive ion.  
(comes from a metal, metal oxide, metal carbonate or metal hydroxide.)

Salts also contain a negative ion.  
(comes from the acid.)

Example:  $\text{CuSO}_4$

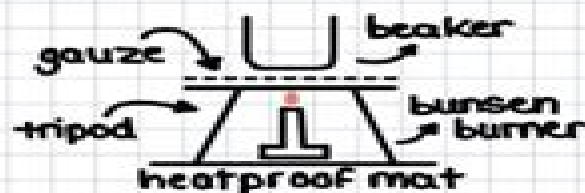
positive ion =  $\text{Cu}^{2+}$

- copper
- copper oxide
- copper carbonate
- copper hydroxide

negative ion =  $\text{SO}_4^{2-}$

- sulphuric acid

### Equipment:



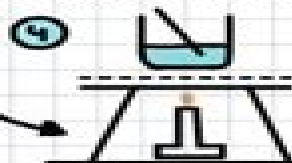
### Method:

Start with a fixed volume of dilute sulphuric acid. This is the limiting reactant.

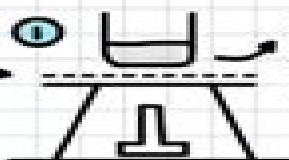
② Gently heat the acid until it is almost boiling.

Then use a spatula to add copper oxide.

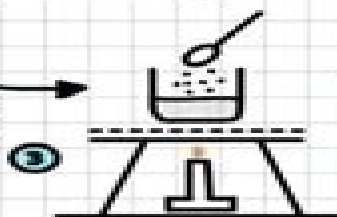
Then stir the solution with a glass rod.



The copper oxide will react and seem to disappear. It will also turn a blue colour.



Fixed volume of sulphuric acid



## REQUIRED PRACTICALS:

C4 - Making salts

C4 - Neutralisation

C4 - Electrolysis

C5 - Temperature changes

## What are salts?

copper sulphate crystals:

blue in colour



⑤

continue adding copper oxide until it is in excess. Stop adding copper oxide if some powder remains after stirring.

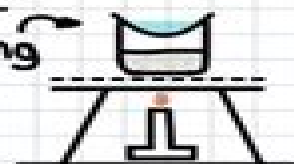
At this point the reaction has stopped.

The solution is now neutral.

Filter out remaining copper oxide.

use an evaporating basin on the tripod to get the copper sulphate crystals.

evaporating basin



# Wednesday 18th January 2012 Chemistry Gcse Paper

**Lei Shi**



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