



Periodic Table of the Elements

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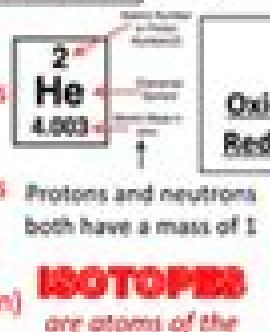
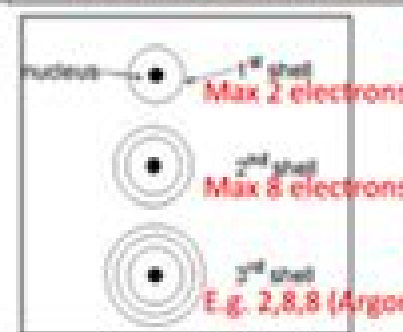
- alkali metals
- alkali earth metals
- transition metals
- post-transition metals
- metalloids
- non-metals
- halogens
- noble gases

An atom:  
Number of Protons = Number of Electrons  
Mass number = Protons + Neutrons

# C2 Chemistry

Periodic table arranges elements by **ATOMIC NUMBER** (proton number/the small one)

The number of outer shell electrons match the group the element is found in.  
E.g. Lithium 2,1 is a group 1 element.

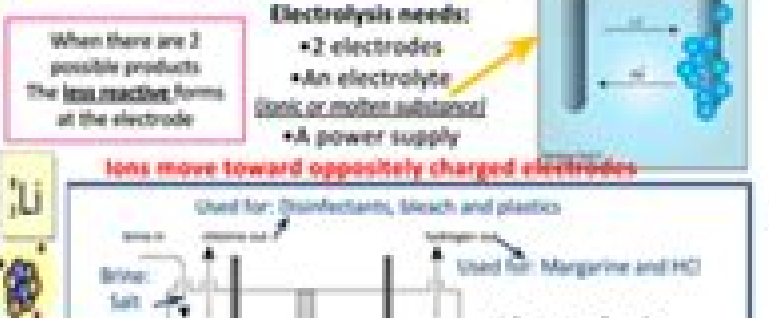
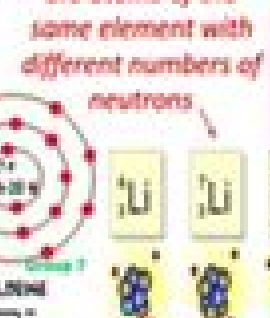
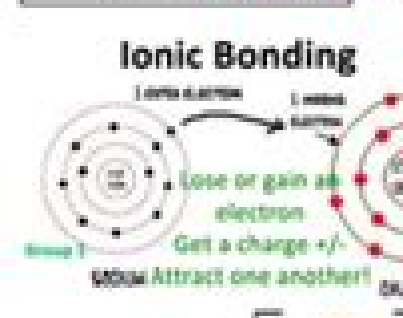
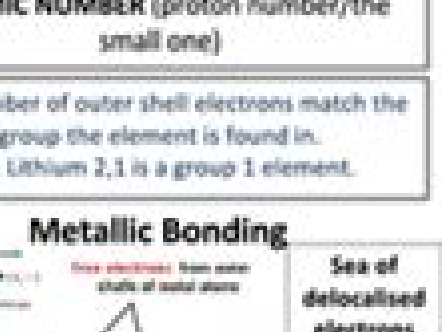
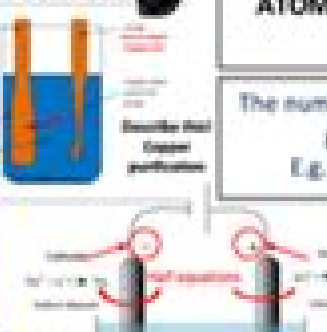


**Electrolysis:**  
Splitting up a substance using electricity

Electrolysis needs:

- 2 electrodes
- An electrolyte (ionic or molten substance)
- A power supply

When there are 2 possible products The less reactive forms at the electrode



Metal ions only arranged in layers  
**ELECTROSTATIC ATTRACTION** between +IONS and ELECTRONS

Metal ions slide over one another making them **MALLEABLE**

Metals conduct **HEAT** well



As a solid: High melting and Boiling Point  
Conducts electricity when melted or dissolved: ions move freely and carry a charge

**NANO-TECHNOLOGY**  
Means **REALLY REALLY REALLY** small!!!!

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**Sam Clegg**



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